

Salen complexes as fire protective agents for thermoplastic polyurethane: Deep EPR spectroscopy investigation

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Abstract:

The contribution to flame retard thermoplastic polyurethane (TPU) of Copper complexes of *Salen* based Schiff bases *N,N'*-bis(salicylidene)ethylenediamine (**C1**), *N,N'*-bis(4-hydroxysalicylidene)ethylenediamine (**C2**) and *N,N'*-bis(5-hydroxysalicylidene)ethylenediamine (**C3**) is investigated in the context of minimizing TPU inherent flammability. Thermal and fire properties of TPU are evaluated. It is observed that fire performances vary depending upon the substitution of *Salen* framework. Cone calorimetry (MLC) results show that, in TPU at 10 wt.-% loading, C2 and C3 reduce peak of heat release (pHRR) by 46 and 50% respectively. At high temperature, these copper complexes undergo polycondensation leading to resorcinol type resin in the condensed phase and thus acting as intumescence reinforcing agents. C3 in TPU is particularly interesting since it delays significantly the time to ignition (MLC experiment). In addition, Pyrolysis combustion flow calorimetry (PCFC) shows reduction in HRR curve suggesting its involvement in gas phase action. Structural changes of copper complexes and radical formation during thermal treatment as well as their influence on fire retardancy of TPU in the condensed phase is investigated by spectroscopic studies supported by microscopic and powder diffraction studies. Electron paramagnetic resonance spectroscopy (EPR) was fully used to follow redox changes of Cu(II) ions as well as radical formation of copper complexes/TPU formulations in their degradation

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3 pathways. Pulsed EPR technique of Hyperfine Sublevel Correlation Spectroscopy
4 (HYSCORE) reveals evolution of local surrounding of copper and radicals with strong
5 contribution of nitrogen fragments in the degradation products. Further, spin state of radicals
6 were investigated by 2D technique of Phase-inverted Echo-Amplitude detected Nutation
7 (PEANUT) experiment. Two different radicals were detected i.e. one mono carbon radical and
8 an oxygen biradical. Thus, the EPR study permits to deeply investigate the mode of action of
9 copper salen complexes in TPU.
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18 1. Introduction:

19 Inorganic products¹ occupy major share in the flame retardant (FR) solution package for
20 engineering polymers,¹⁻¹³ elastomers,¹⁴⁻¹⁵ insulation materials/coatings and textiles.¹⁶ Metals
21 across the periodic table were studied as FR since they can occur in diverse forms (e.g. oxides,
22 hydroxides, carbonates, and other salts) and they can virtually offer unlimited possibilities.
23 Metal bound compounds provide effective flame retardancy by several routes. Their role in FR
24 solution is manifold although their primarily used as a FR additive or a synergist.¹⁴ They can
25 act in both condensed and gas phases. They are also considered for their role as
26 compatibilizers/processing aid, smoke suppressants, stabilizers, char reinforcing agents, char
27 former, ceramization/glass former, catalyst, and cross-linking agents irrespective of their
28 hierarchical level at their molecular organization. They may have endothermic effect by
29 absorbing heat from combustion zone. Hydrated metal compounds releases water at higher
30 temperature which acts as fuel diluent. In case of magnesium hydroxide ($Mg(OH)_2$) the
31 dehydration temperature is superior to 300°C which is unusually higher unlike classical
32 hydrated compounds which are normally released around 100-150°C. Some metal oxides
33 formed during burning migrate and reinforce char surface that can deflect heat and acts as a
34 thermal shield.
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48 Simple metal salts (e.g. Mg^{2+} , Al^{3+} , Ti^{4+} , Lanthanides) and oxides are the basis of this database
49 of inorganic functionalities.¹⁻⁸ Salts of higher compositional order/chelates such as metal
50 acetates, acetonates, oxalates, borates, phosphates, phosphinates, double salts (e.g. Alum), are
51 also well known in FR formulations. Next candidates in this level of molecular hierarchy are
52 inorganic-organic hybrid molecules that have emerged as contemporary multi-functional
53 additives.^{17,18} Melamine intercalated metal polyphosphates (e.g. Safire[®]) belongs to this
54 category. Coordination compounds or metal complexes were relatively new in this domain of
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FR whose admission was restricted largely due to the misunderstanding of their limited thermal stability.¹⁵ Silicates, zeolites, organically modified clays were already extensively used in this domain.^{19,20} The other end of this spectrum of molecular hierarchy is tailorable coordination polymers (CP)/metal-organic frameworks (MOF) which are unexplored in the scope of FR (figure 1).²¹

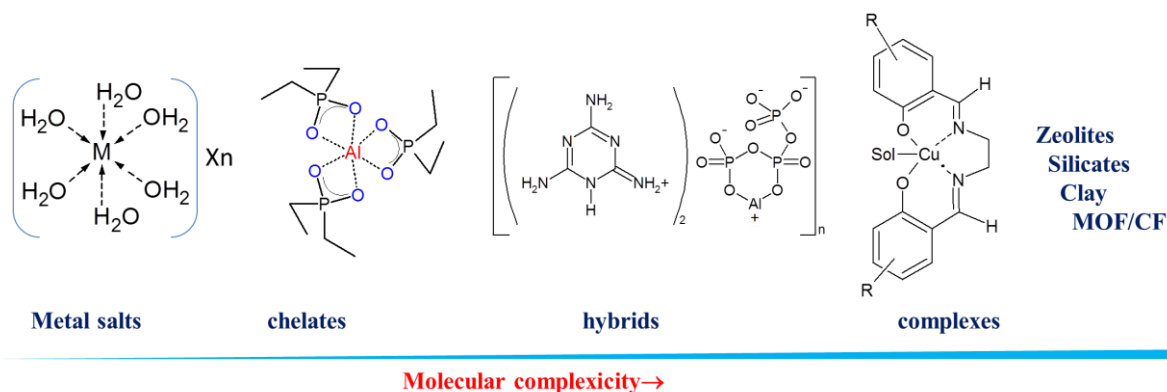


Figure 1. Tentative classification of inorganic elements used in flame retardancy (X=anion, n=number, Sol=solvent, R= substituant, e.g. hydroxyl group) (MOF: metal-organic frameworks; CP: coordination polymer).

Coordination complexes or simply metal complexes are molecular entities wherein pre-designed organic molecules (also called ligands) are allowed to react with metal salts to have topologically diverse molecular architecture combing the properties of both organic and inorganic functionalities. They can offer wide range of tailorability to target application with the help of molecular engineering concept. Their distribution, organization and properties in the host matrix (e.g. polymers) depends on several factors like polarity of molecules, functional groups on the periphery of the molecules, reactivity of metal centers, unsaturation at metal centers, supramolecular interactions, coordination number of metal centers, thermal stability, hydrophobic/philic interactions etc. Metal salts or simple metal chelates have high tendency to form coordinate bonds with wide range of donor atoms of host matrices often changing the physical properties. On the other hand, metal complexes may limit the interaction of metal ions in the host matrix and if their demetallation takes place, their interference with functional groups of matrix is restricted. However, their high temperature applications are still in their infancy.

A recent work by Fontaine et. al on copper complexes based on ‘Salen’ molecules in the flame retardancy of thermoplastic polyurethane revitalized the concept of ‘tailored’ metal complexes

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3 for high temperature protective applications.¹⁵ Although metal bound additives were
4 extensively used in the flame retardancy of polypropylene, polyethylene, PVC, TPU, epoxy,
5 polyamides, polycarbonates, EVA, PVA, cotton fibers etc.¹⁻¹⁴ the unique feature about this
6 *Salen* copper complexes were their simple molecular framework. Mostly such designs were
7 used in molecular sensors, catalysis, and biological applications but been ignored for high
8 temperature applications.^{22,23} In our recent study we have shown how such well-designed
9 organic molecules could undergo *in situ* structural transformation forming thermally stable
10 resins and improve fire weakness of TPU.²²

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12 Present work is about elucidating the role of such *Salen* copper complexes that is an extension
13 of our previous work¹⁵ in the fire protection of TPU. The reason for choosing TPU is its
14 impressive range of performance characteristics, and high demand for their fire protection
15 requirements to extend their service life in coatings, adhesives, flexible foams and other
16 applications.^{24,25} We have used substituted and un-substituted *Salen* molecules for
17 complexation with copper salts. Fire performance, thermal properties, characterization of
18 complexes in the TPU matrix and their degradation pathway leading to fire protection are
19 discussed.

2. Experimental

2.1. Materials

20 Thermoplastic Polyurethane (TPU) C85A (Elastollan) polyester polyurethane) was kindly
21 supplied by BASF.²⁹ Copper complexes of *N,N'*-bis(salicylidene)ethylenediamine (C1) and
22 *N,N'*-bis(4-hydroxysalicylidene)ethylenediamine (C2) and *N,N'*-bis(5-
23 hydroxysalicylidene)ethylenediamine (C3) were synthesized based on earlier reports.¹⁵

2.2. Formulations, processing, and sampling

24 TPU is dried for 12 hours at 80°C before use. Compounding of formulations with 10 wt % of
25 C1, C2 or C3 were performed using HAAKE Rheomix OS PTW 16 blender.¹⁵ The temperature
26 of the mixer was set at 180°C and the shear was 50 rpm for 10 minutes. These blended mixtures
27 were ground in liquid nitrogen in an ultracentrifuge mill to produce a powder and dried at 80
28 °C under vacuum for 12 h before use.

29 Samples for elucidating fire protection and decomposition pathways were mainly obtained
30 from thermogravimetric studies (TGA) wherein TGA at the end of each degradation step was
31 stopped and sample were collected. Samples were also collected from cone calorimetry

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3 experiment which is based on HRR (heat release rate) curve wherein sample (100 x 100 x 3
4 mm³ plates) exposure to heat flux were stopped and withdrawn at the onset, peak, descending,
5 and end of HRR curve. Onset, peak and residue in the manuscript means degradation samples
6 collected at the beginning, peak and end of the MLC experiment. A 4x4cm piece is scooped
7 out from the central part of these plates including the intumescence part. Residue are manually
8 crushed and collected. Other samples are semi-solid and are cut into small pieces and ground
9 in liquid nitrogen in an ultracentrifuge mill to produce a powder. Samples were dried at 100 °C
10 under vacuum for 12 h and stored. These samples are used for some analysis.
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19 2.3. Instrumental

20 2.3.1. Thermal analysis:

21 Thermogravimetric analyses (TGA) were performed using SDT Q600 (TA instruments).
22 Samples (approx. 6-7 mg) were placed in open alumina pans and heated under nitrogen
23 atmosphere (100 mL/min) with a heating rate of 10°C/min.
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29 2.3.2. Mass loss cone calorimeter:

30 The mass loss cone calorimeter (Fire Testing Technology (FTT), ISO 13927, ASTM E906) is
31 used for recording heat release rate (HRR) curve. Plates (100x100x3 mm³ plates) for cone
32 calorimeter test were made via compression molding using DARRAGON press apparatus (50
33 kN, 6min, 185°C). Plates were wrapped in aluminium foil leaving the upper surface exposed
34 to the heater and placed in horizontal position on ceramic block encased in a metallic container
35 at a distance of 35mm from cone base. External heat flux of 35kW/m² was used for all the
36 experiments. The experiments were done in triplicate and the error on the measured parameters
37 is within 10%.
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46 2.3.3. Pyrolysis combustion flow calorimetry (PCFC):

47 Fire performance was evaluated based on heat release rate (HRR) curve obtained from
48 FAA Micro calorimeter (FAA Fire testing technology) operated at 1°C/s to 700°C in
49 the pyrolysis zone according to ASTM D7309-07 method A. The combustion zone was
50 set at 900°C. Oxygen and nitrogen flow rates were set at 20 and 80 cc/min respectively.
51 Repeatability is confirmed by three consecutive trials.
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2.3.4. Spectral and microscopy analysis:

FT-IR were recorded on Nicolet iS 50 in ATR mode at room temperature. VHX digital optical microscope (Keyence, VH-Z 100R) was used to investigate texture of samples. The surface of the sample for EPMA was prepared with a Leica Ultracut ultramicrotome at cryo temperature with a diatome diamond knife. The images were taken with a Cameca SX100 EPMA in BSE mode at 15kV 15 nA. Morphology of char and some residues were studied using scanning electron microscope (SEM) (Field emission gun (FEG) Hitachi S4700).

CW EPR/Pulsed EPR: EPR continuous-wave (CW) experiments were performed on X-Band Bruker ELEXSYS E500 spectrometer operating at 9.78 GHz. The powder was inserted in quartz tube and the spectra were recorded on TM cavity at room temperature under non-saturated conditions: microwave power in a range of 1-12 mW and modulation amplitude in a range of 2-6 G. High temperature experiments were recorded on ER4114HT cavity which was purged by nitrogen. The powder was inserted in Tailor-made sample tubes. The temperature was measured with a thermocouple (type K: Cr-Al) close to the sample. The gas flow rate and the temperature are controlled with the temperature unit (ER 4131VT). The heating thermal power is provided by the power booster ER4131VT-1020.

Pulsed EPR was studied at 5K with a Bruker ELEXSYS E580 spectrometer equipped with a helium flow cryostat. 2-pulses echo field sweep acquisitions were performed using standard Hahn echo sequence $90^\circ\text{-}\tau\text{-}180^\circ\text{-}$ with 90° pulse length of 16 ns and τ value of 136 ns. The HYperfine Sublevel CORrElation spectroscopy (HYSCORE)²⁶ experiments were recorded with 256×256 data points for both t_1 and t_2 time domains, 90° pulse length of 10 or 24 ns (respectively for copper and radicals) and an echo delay of 136 ns. The obtained HYSCORE spectra are composed of two quadrants: the first quadrant (+,-) where $A > 2\nu_I$ (ν_I being the nuclear frequency) corresponding to strong hyperfine coupling A between the I nucleus and the unpaired electron and the second quadrant (+,+) where $A < 2\nu_I$ corresponding weaker interactions. Direct measurement of the spin (S) was performed by Phase-inverted Echo-Amplitude detected Nutation (PEANUT) experiment²⁷ using a two dimensional versus field sweep. The non-selective $\pi/2$ preparation pulse of length t_p of 16 ns was performed with two steps phase cycling (+X, -X). The spectrum was recorded with 256×256 data points along both dimensions, 90° pulse length of 16 ns separated from the nutation pulse by the time interval of 136 ns. Spectra are recorded at 5K.

2.3.5. TGA-FTIR analysis:

Gas phase analysis were carried out in TGA Q5000 (TA instruments) coupled with FTIR Nicolet spectrometer (ThermoFischer). Samples (~10mg) were heated in a 250 μ L alumina crucible from 50°C to 800°C with a heating rate of 10°C/min under nitrogen atmosphere. A balance purge flow of 15 mL/min and a sample purge flow of 100mL/min was maintained. A transfer line with an inner diameter of 1mm was used to connect TGA and infrared cell. The temperature of transfer line and gas cell was kept at 225°C. Prior to this, samples were kept for 2h under nitrogen stream. IR spectra were collected in 400-4000 cm^{-1} spectral range.

2.3.6. Pyrolysis GC/MS:

Samples (~200 μ g) were analyzed by Pyrolysis GC/MS (Shimadzu, GCMS-QP2010 SE). GC separation was carried out with a fused silica capillary column (SLB 5ms) of 30m length and 0.25 μ m thickness. Analyses were carried out both in direct pyrolysis mode and desorption method. The temperature selection for desorption is based TGA pattern of concerned sample with a heating rate of 10°C/min. The initial column temperature was held at 35 °C for a period that corresponds to desorption time followed by a temperature ramp at 10°C/min to a final temperature of 300 °C and isotherm for 20 min. For direct pyrolysis the furnace is set for the final temperature (600°C which corresponds to completion of second major degradation step) and sample is pyrolysed for 0.5 min. Column oven temperature is programmed in the following way. The initial column temperature was held at 35 °C for 1 min. followed by a temperature ramp at 10°C/min to a final temperature of 300°C and isotherm for 20 min. Helium was used as a carrier gas at pressure of 120kPa with a split ratio of 50. The transfer line was maintained at 275°C. The MS was operated under Electron Ionization EI mode. An online computer using GCMS real time analysis and PY-2020i software controlled GC/MS system. The eluted components were identified by library search and only significant peaks (above 90% match) observed in the total ion chromatograms were studied and compared to a mass spectral database (GCMS postrun analysis, and NIST).

3. Results and discussion

3.1. Fire performances

Flammability characteristic of TPU and flame retarded formulations (with 10wt % of C1-C3) are evaluated from heat release rate (HRR) curves derived from two different techniques -

Mass loss calorimetry (MLC) and Pyrolysis combustion flow calorimetry (PCFC). They are displayed in Figure 2a and 2b and parameters are given in Table 1.

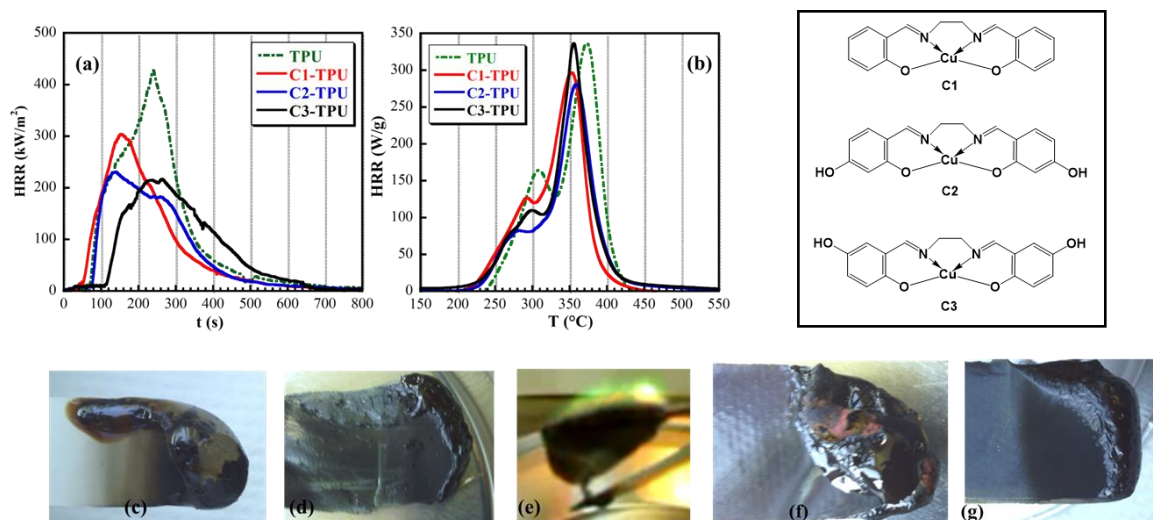


Figure 2. (a) HRR curve obtained from mass cone calorimetry. (b) HRR curve obtained from PCFC. (c, d, f, g) Response to flaming under LOI condition (c) neat TPU (d) C1-TPU (e) A piece of C1-TPU burning with green flame (f) Intumescence in C2-TPU also showing red patches (g) Slight intumescence in C3-TPU.

Neat TPU under cone calorimetry conditions burns almost totally resulting in small amount of brittle char (pHRR, 426 kW/m²). C1-TPU shows 30% reduction in pHRR and C2-TPU shows 46% reduction of the pHRR with a shoulder on the pHRR. C3-TPU exhibits the best behaviour with nearly 50% reduction in pHRR and stretches the ignition (sustain) for longer time (110s vs. 65 s for neat TPU). Total heat release (THR) for TPU is 83.8 MJ/m² and decreases to 60.9 MJ/m² for C1 and is nearly in the same range for C2-TPU and C3-TPU i.e. 56.8 MJ/m² and 59.5 MJ/m² respectively (Table 1). Unlike TPU, C1-TPU and C2-TPU time to ignition (TTI) for C3-TPU is highest with 110s against values range of 65-75 s. PCFC of neat TPU shows two peaks with the first one whose being smaller. Both peaks are slightly reduced in the case of C1-TPU formulation. Interestingly C2-TPU shows reduction in HRR in both peaks that is considerable in the first step (50% reduction in pHRR) compared to C1 and C3. C3-TPU does not show significant reduction in these values. In all the formulations, both signals are slightly shifted to lower temperature. LOI test did not show any improvement, the LOI values range is 23-24 vol-% (Table 1). While C1-TPU displays dripping, C2-TPU shows intumescence whereas C3-TPU shows both dripping and intumescence (Figure 2).

Table 1. Fire properties derived from mass loss calorimeter, PCFC and LOI for TPU and formulations.

Materials	Mass loss calorimeter parameters			PCFC parameters		LOI (%)
	Time to ignition (s)	pHRR (kW/m ²)	THR (MJ/m ²)	pHRR (W/g)	THR (kJ/g)	
TPU	65	426	83.8	164, 336	25.7	23
C1-TPU	50	300 (-30%)	60.9 (-27%)	126, 295	22.8	23
C2-TPU	75	229 (-46%)	56.8(-32%)	81, 280	20.9	23
C3-TPU	110	217 (-49%)	59.5(-29%)	77, 110, 335	22.6	24

In the following sections each complex is discussed separately focussing their thermal properties, distribution in TPU matrix and their possible role in thermal protection.

3.2. Case study of C1-TPU

3.2.1. C1 in TPU matrix

C1 is a green microcrystalline solid and upon blending in TPU imparts green colour to the TPU. Reported three-dimensional X-ray methods shows that C1 is a four coordinate, mononuclear complex and the geometry varies with solvent of crystallization.²⁸ A section in supplementary information is dedicated to the preliminary characterisation and thermal properties of C1 (Figure S1-S4). BSE image and elemental mapping (Figure 3) by EPMA studies have shown that C1 was uniformly distributed in TPU matrix.

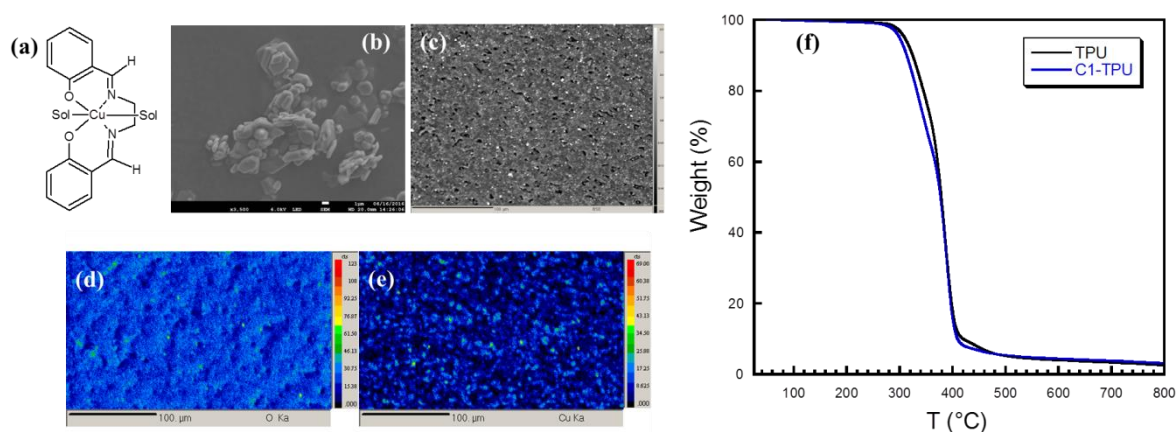


Figure 3. (a) Tentative molecular structure of C1. (b) Microscopic image on C1. (c) BSE image on C1-TPU. (d,e) Elemental mapping (oxygen and copper respectively) by EPMA showing distribution of C1 in C1-TPU. (f) TGA of TPU and C1-TPU (under nitrogen atmosphere and 10°C/min).

By comparing FTIR of C1-TPU (Figure S4) and neat TPU it is found that none of the TPU signals are affected ($\nu(\text{NH})$ (3331, 3305 cm^{-1} doublet) and $\nu(\text{CH})$ signal (sharp signals at 2922 and 2852 cm^{-1})). The 1704 cm^{-1} band due to $-\text{C}=\text{O}$ stretching vibration of urethane and $-\text{C}=\text{O}$ stretch of polyol signal at 1727 cm^{-1} are also unaffected.³⁰⁻³³ When compared to C1 additional signals are observed corresponding to main characteristic bands of C1.

3.2.2. Thermal degradation and associated structural changes

In TGA studies (Figure 3f, Table 2) it is found that neat TPU begin to lose weight slowly above 250°C. Although there is no clear distinction between decomposition steps, weight loss in the first step is gradual until 360°C and thereafter is more pronounced until 430°C. Small amount of residue is left after 500°C (i.e. 2.5 % at 800°C).

Table 2. Thermal properties of TPU and formulations (under nitrogen atmosphere and 10°C/min).

Material	1st step (T, °C) (weight loss, %)	2 nd step (T, °C) (weight loss, %)	Residue at 500°C
TPU	50-350 (22.5)	350-430 (68.3)	2.9
C1-TPU	50-357 (34.4)	357-429 (58%)	3.1
C2-TPU	50-339 (16.2)	339-463 (68.4)	10.5
C3-TPU	50-359 (25.1)	359-443 (59.7)	9.0

TG curve of neat C1 shows that it is stable until 323°C that is quite high for a Schiff base complex (Figure S1). Decomposition curve of C1-TPU took deviation quite early in the first gradual step (compared to TPU) indicating a slight destabilising effect but behaves much similarly to TPU in the second step and also in terms of residue (3.1 %). The residue amount indicates that C1 could not bring out significant changes to promote char formation (as also seen in LOI experiment).

Samples at the end of first stage of decomposition is collected and analysed by FTIR. C1-TPU-358 (sample collected at 358°C from TGA experiment) which represents the end of the first step of decomposition shows that 3346 cm^{-1} band due to $-\text{NH}$ disappears. 1704 cm^{-1} band due

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3 to -C=O stretching vibration of urethane is reduced that indicates breakdown of urethane group.
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5 However, the -C=O stretch of polyol signal (1727 cm^{-1}) is unaffected. These changes were
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7 also observed in the degradation of neat TPU. Along with this, band at 1528 cm^{-1} (semicircle
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9 aromatic stretch) decreases in intensity and 1223 cm^{-1} signal due to -C-O stretch is disappeared.
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11 If we compare FTIR of TPU-348 and C1-TPU-358, we can see that additional bands are
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13 observed at $1617, 1572, 757\text{ cm}^{-1}$ which indicates formation of some new aromatic structure
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15 (Figure S5). No further analysis was done on small amount of char at the end of analysis.

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17 DSC profile of C1-TPU is much similar to TPU (Table S2) with broad endothermic signal
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19 around 175°C due to disordering of hard segment (HS) crystallites and at 214°C corresponds
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21 to melting of HS of TPU. In the cooling mode an exothermic peak is observed (83°C) due to
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23 crystallization.

24 3.2.3. Discussion on C1 Performances in TPU

25 Based on the above results and on structural changes upon thermal treatment the effect of the
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27 C1 copper complex on fire retardancy of TPU is followed. This is mostly focussed on the
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29 analysis of residues of C1-TPU formulation collected either from TGA or from MLC
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31 experiment. The aromatic species found in the residue of C1-TPU-358 is from degrading
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33 fragments of C1, possibly containing phenolic group bonded to copper. But this new aromatic
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35 species could not promote a strong char formation and is unable to hold the molten polymer.
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37 Under MLC condition, before ignition it is observed that C1-TPU initially displays weak
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39 intumescence, which soon collapse. Interestingly during cone test flames were imparted with
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41 green colour. It is due to some part of sublimation of complex (or free copper ions from Salen
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43 framework after de-metallation) and subsequent atomisation of copper ions that gives
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45 characteristic green colour to the flame. Thus significant amount of copper ions are lost and
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47 possible action of C1 in condensed phase is lost. During py-GC/MS experiments (on C1)
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49 sublimation of *Salen* ligand is detected that shows that ligand part is dissociated from copper
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51 ions during the thermal treatment and escapes as volatiles (although its detection is below the
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53 expected level in py-GC/MS of C1-TPU). Further, no additional char is observed from TGA.
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55 Thus there is no significant contribution of C1 or copper in the condensed phase. PCFC results
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57 show little reduction of the pHRR and THR, therefore C1 present a little contribution in the
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59 gas phase. The observed reduction in MLC it is maybe also due to small amount of char swelled
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61 by volatiles accompanied.

3.3. Case study of C2-TPU

3.3.1. C2 in TPU matrix

Regarding the MLC test, it appears that the reduction's peak HRR (46%) is more significant in C2-TPU than its non-hydroxy analogue (C1) making it worth to explore in detail. C2-TPU is a violet microcrystalline solid. EPMA shows the distribution of additives in the TPU matrix (Figure 4b-d). Due to large particle size, dispersion seems to be less than for C1-TPU. TPU matrix seems to be unaffected by the insertion of C2 in TPU as shown by FTIR (comparison of C2-TPU and C2, Figure S6). The band at 1620 cm^{-1} in C2 is clearly seen in C2-TPU but most of other signals are overlapped. The $\nu(\text{OH})$ of C2 at 3442 cm^{-1} is not seen which may be due to its involvement in hydrogen bonding with the matrix. Another possibility is the structural changes of C2 upon insertion in TPU matrix. This could be due to loss of coordinated and non-coordinated water molecules. This is highly probable because the processing temperature of C2 in TPU is around 180°C which is high enough to remove the solvent molecules. It can also be confirmed by TGA of neat C2 (Figure S1) which display small initial weight loss step until 100°C . Further structural characterization of C2-TPU by solid state NMR as we reported for Salen-TPU²² was not possible due to presence of paramagnetic metal ions. However, a complementary technique of Electron Paramagnetic Resonance spectroscopy (EPR) was used to explore structural and electronic changes due to copper ions in C2. To confirm the redox state of the copper, continuous wave-EPR experiments were performed on C2 and C2-TPU collected after thermal treatment. Spectra of C2 and C2-TPU presented in figure 4e, display a broad signal attributable copper (II) (d^9 , $S=1/2$) state. The different 'g' factor values obtained and line-shapes are significantly different which suggest the changes of copper coordination after insertion in TPU matrix as it was observed by FTIR.

3.3.2. Thermal degradation and associated structural changes

Thermal stability of C2-TPU is different from that of C1-TPU (Figure 4f). It begins to lose weight around 263°C similarly to neat TPU. The first step of decomposition until 341°C is less gradual than neat TPU. The second step is shorter than the TPU one with higher residue weight at the end of the step. Residue weight at 500°C (11.5%) is quite significant compared to neat C1-TPU (3.1%).

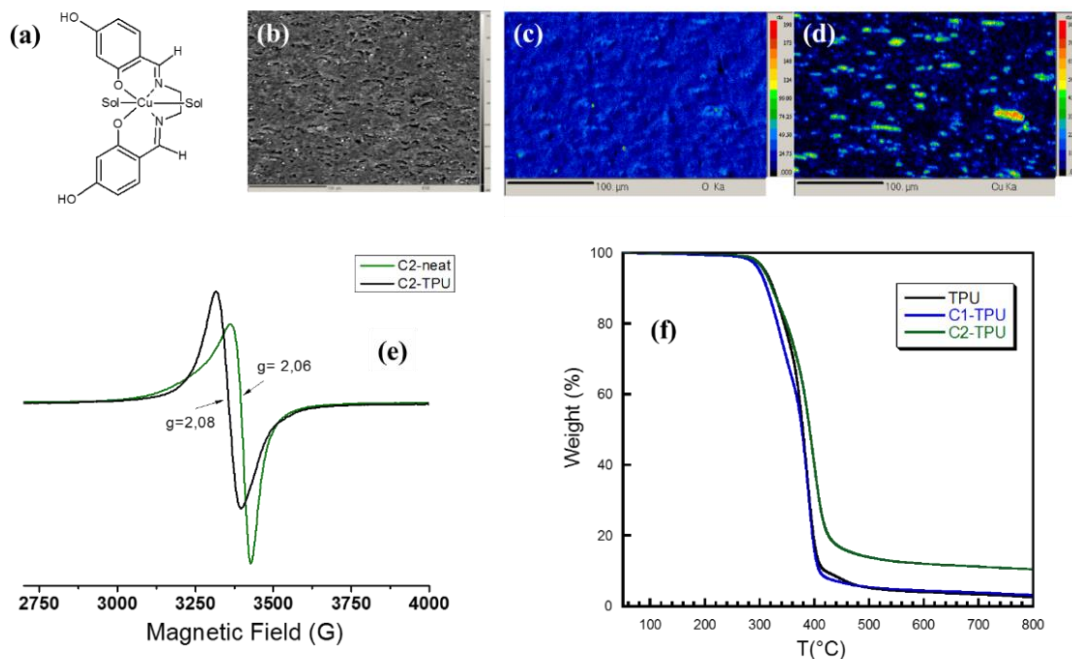
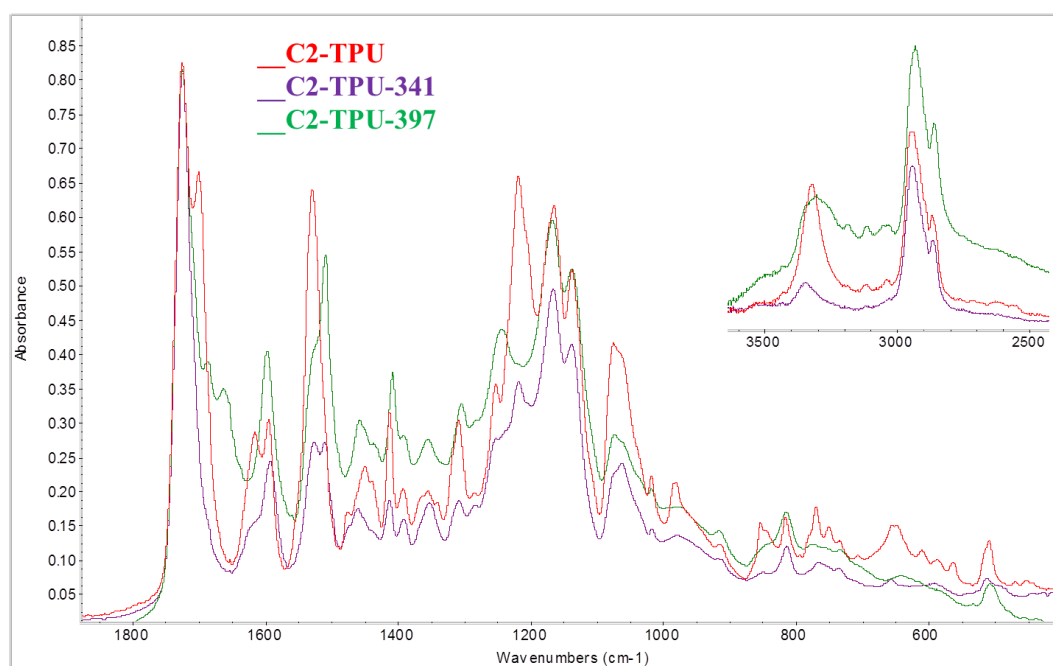


Figure 4. (a) Tentative structural formula of C2. (b) BSE image on C2-TPU (c, d) oxygen and copper mapping respectively in C2-TPU. (e) CW-EPR spectra on neat C2 and C2-TPU. (f) TGA of TPU, C1-TPU and C2-TPU (nitrogen atmosphere, 10°C/min).

Thermal degradation investigation of C2-TPU is based on sample collected from TGA. TGA was stopped at the end of each degradation step (Table 2) and samples were analysed by FTIR. C2-TPU loses weight (16%) until 341°C which constitutes first major weight loss step. Analysis on the residue shows that in C2-TPU-341 (which is a stage after initial minor weight loss step) the band of urethane ($\nu(\text{C}=\text{O})$) has disappeared, and NH stretch (3349 cm^{-1}) decreased also in intensity. This denotes break down of TPU chain. The band at 1619 cm^{-1} due to imine in C2 also decreases. It is also observed a decrease of aromatic stretch at 1532 cm^{-1} and $\text{C}-\text{O}$ stretch at 1220 cm^{-1} in C2. This means that C2 has undergone structural changes during the first TGA degradation step. The degradation product of the first step of decomposition was analysed by TGA-FTIR. From 255 °C (onset of 1st decomposition) hydrocarbon fragments, CO_2 , isocyanate (weak signals) are detected. Around 268°C, the evolution of hydrocarbon ($2998, 2874\text{ cm}^{-1}$), small amount of CO_2 ($2371, 2308, 670\text{ cm}^{-1}$), species containing $\text{C}-\text{OH}$ and $\text{C}-\text{O}-\text{C}$ group ($1079, 912\text{ cm}^{-1}$) and hydroxyl species (bands above 3594 cm^{-1}) are observed. Based on the reported data and as seen in neat TPU this latter product is assigned to butanediol which is also confirmed from py-GC/MS test (Table S3). Second step of decomposition runs between 341 to 463°C. TGA-FTIR shows that until end of second step of decomposition large

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3 quantity of hydrocarbon, ketone (1766 cm^{-1}) and small quantity of CO_2 , and other ester
4 fragments of TPU ($1417, 1142, 1070, 922\text{ cm}^{-1}$) are detected.
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8 Residue left after the second step is analysed by FTIR (Figure 5). C2-TPU-397, which is end
9 of decomposition, shows several spectral changes. There are several additional bands around
10 $1511, 1600\text{-}1665\text{ cm}^{-1}$ region characteristic of aromatic species. This corresponds to new
11 substituted aromatic species or aromatic groups from degrading TPU matrix. The residue
12 remaining in TPU at 400°C is $\sim 5\%$ whereas it is 14% in C2-TPU that shows significant amount
13 of char formation. We have seen from MLC that C2-TPU witness intumescence during fire test
14 (Figure S9) and also during LOI. Regarding MLC experiment, this residue formation in the
15 form of char could help to reduce the pHRR and moreover, unlike other samples, C2-TPU
16 behaviour is shared over two bands in HRR curve.^{34b}
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46 **Figure 5.** Tracking structural changes of C2-TPU formulation by FTIR
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51 It is also found that neat C2 also produce significant amount of residue (46%) in TGA analysis
52 at 500°C (Section S1).
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55 3.2.3. Discussion on C2 performances in TPU 56

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58 Considering the potential of undergoing resin formation of its precursor (L2)^{22,23} the higher
59 residue weight found in C2 hints to some sort of structural transformation leading to stable
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3 polymeric structure, promoting char. Temperature treatment in C2 although does not lead to
4 complete turnover to phenolic/resorcinol resin type material as seen in L2, there is an evidence
5 of structural transformation promoting char formation.³¹ After initial thermochromic switch
6 (Figure S7) in C2 the degradation after 400°C is very gradual and extends until 800°C (Figure
7 S1). This process of poly-condensation and resin formation is also expected during C2-TPU
8 degradation. However, this process gets complicated due to the presence of copper ions and its
9 structural and redox chemistry. This will be tedious to follow in TPU due to the interference
10 from degrading macromolecules and its tendency to unzipping in multiple way at higher
11 temperature. There is evidence from FTIR that some TPU fragments are retained in the residue
12 and thus being a source for char formation. It is possible that char formation have contribution
13 from deviation in unzipping TPU matrix influenced by C2 and also structural transformation
14 of C2 during thermal treatment. Thus, C2-TPU works mainly by char formation.

3.4. Case study of C3-TPU:

25 C3-TPU shows flame retardancy by intumescence that is slightly different from C2-TPU. This
26 intumescent char is soft and flexible and develops from highly viscous molten intermediate.
27 The significant reduction in pHRR with delayed time to ignition makes C3 a flame retardant
28 for TPU. Under LOI condition, a slight intumescence is observed and the molten state also
29 contributes in the form of dripping

3.4.1. C3 in TPU matrix:

30 C3 is brownish yellow crystalline (Figure 6e) compound that imparts beige colour to C3-TPU.
31 EPMA (Figure 6f) shows the distribution of C3 in TPU matrix, unlike C2, bigger particles or
32 aggregates of complex are observed in the matrix. TPU matrix seems to be unaffected by the
33 copper complex as evidenced by FTIR of C3-TPU (Figure S10). However, the IR bands
34 attributed to C3 have undergone some changes. Upon insertion into TPU, the broad band at
35 3164 cm⁻¹ due to -OH is not seen in C3-TPU that is due to its involvement in hydrogen bonding
36 with the matrix. The $\nu(\text{C}=\text{N})$ and aromatic ring stretches in C3 (1634 cm⁻¹ s br, 1613 cm⁻¹ sh,
37 1547 cm⁻¹ s, 1472 cm⁻¹ m) are also slightly shifted and appeared with lower intensity. The
38 phenolic $\nu(\text{CO})$ that appears at 1267 cm⁻¹ and 1159 cm⁻¹ in C3 is seen but with lower intensity.
39 This suggest that C3 upon distribution TPU matrix did not retain its original geometry. This is
40 due to loss of coordinated water molecules during the processing with TPU that changes the
41 geometry around copper and re-alignment of hydrogen bonding with the matrix. This is further
42 supported by the change in EPR signal of C3 in C3-TPU (figure 6b and c). As observed on C2-
43 TPU samples, CW-EPR study on C3 and C3-TPU powders indicates changes in molecular
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environment around copper (figure 6b and c). Indeed at room temperature, C3 present a signal of Copper (II) state with an axial symmetry ($g_{\parallel} = 2.18$ and $g_{\perp} = 2.08$) whereas C3-TPU display a broad signal centered at 'g' value of 2.07.

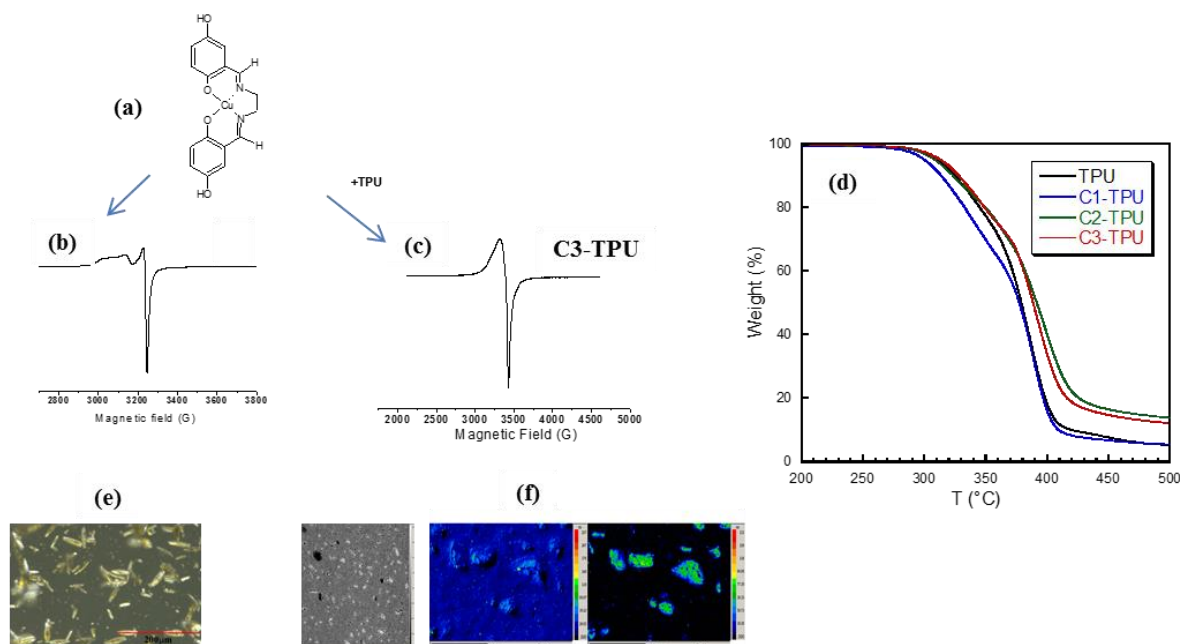
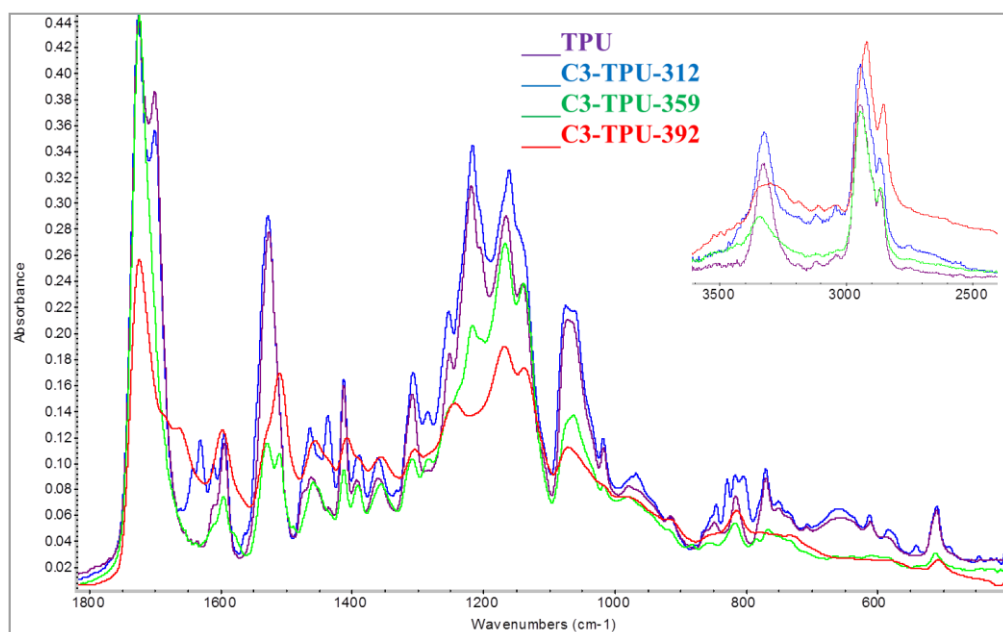


Figure 6. (a) Structure of C3. (b) CW-EPR spectrum of C3 at room temperature. (c) CW-EPR spectrum of C3-TPU at room temperature. (d) TGA profile (e) Crystalline C3 (f) EPMA on C3-TPU showing distribution.

3.4.2. Thermal degradation and associated structural changes

TGA profile of C3-TPU is much similar to C2-TPU (Figure 6d). C3-TPU starts to lose weight from 121°C (slow mass loss rate) until the onset of first step of decomposition (268°C). The first step of decomposition shows loss of 25% weight until 359°C. Volatile degradation products during this step were obtained from *in situ* operando technique of TGA-FTIR and was appropriately supported by py-GC/MS test. TGA-FTIR shows around 268°C, the evolution of hydrocarbon ($2998, 2874 \text{ cm}^{-1}$), small amount of CO_2 (medium intensity of signals) ($2371, 2308, 670 \text{ cm}^{-1}$) species containing $-\text{C}-\text{OH}$ and $\text{C}-\text{O}-\text{C}$ group ($1079, 912 \text{ cm}^{-1}$) and hydroxyl species (bands above 3594 cm^{-1}). Based on the reported data and as seen in neat TPU this is assigned to butanediol which is also confirmed by py-GC/MS test.^{22,25} A carbonyl species (1718 cm^{-1}) is seen around 310°C which could be a fraction containing urethane or ester. Residues were collected at the end of each TGA steps (312°C, 359°C, 396°C) and were analysed by FTIR (Figure 7).

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3 FTIR of C3-TPU-359 (Figure 7) which is a stage after initial minor weight loss step shows that
4 the band of urethane (1702 cm^{-1}) has disappeared as seen in C2-TPU, NH stretch also decreased
5 in intensity (3349 cm^{-1}). It was already seen in case of L2-TPU that it is carbonyl stretching
6 (1704 cm^{-1}) of urethane group that undergoes reduction in intensity confirming our earlier
7 observation that hard segment (HS) of TPU is the preferred choice of cleavage.²² Multiple
8 bands around 1620 cm^{-1} due to C3 which is clearly visible in the spectrum of C3-TPU are now
9 nearly disappeared and also the other prominent bands of C3. This means that C3 in TPU has
10 further undergone severe structural changes by this stage.
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39 **Figure 7.** Structural changes followed by FTIR in the decomposition of C3-TPU.

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42 Second step in TGA is slightly less gradual than C2-TPU and ended by 420°C . The mass loss
43 of volatiles was analysed by TGA coupled FTIR. At the beginning of this second step,
44 hydrocarbon and CO_2 are produced in large quantity along with carbonyl species and diols.
45 There is an additional carbonyl band (1770 cm^{-1}) which is due to cyclopentanone which is a
46 cyclized product formed during the decomposition of TPU. This evolution is supported by py-
47 GC/MS test. There is also contribution from linear ester fragments from soft segment (SS) of
48 TPU which considerably increases as the 2nd step reaches to its maximum weight loss. MDI
49 (diphenyl isocyanate) although not evidenced in FTIR (overlapped by broad and strong CO_2
50 band) was clearly identified in py-GC/MS of neat C3. From py-GC/MS chemical species like
51 but-3-en-1-yl propionate (MW 128), but-3-en-1-yl pentanoate (MW 156), 4-hydroxybutyl
52 pentanoate (MW 174), adipic acid 4-butenyl ester (MW 200), 4-hydroxy butyl adipic acid
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3 monoester (MW 218), dibutyl adipate (MW 254) are identified. There was no significant
4 difference of their production in all formulations. It is highly likely that some of these species
5 containing potential donor atoms may bind to copper ions and remains in the condensed phase.
6 Residue collected around 400°C is studied by FTIR. Interestingly band at 1727 cm^{-1} is retained
7 and is assigned to segmented polyol from TPU. This is expected to be involved in char
8 formation. There are reports of involvement of hydrocarbon fragments in the formation of char
9 due to radical formation.²⁴ There are several additional bands around 1511 cm^{-1} , and between
10 1600-1685 cm^{-1} which is the region characteristic of aromatic species. As the aromatic
11 contributor MDI (from TPU) decompose completely by this was evidenced by py-GC/MS, the
12 observed aromatic species are assumed to originate from C3. It can be recalled that the *Salen*
13 precursor of C3 is capable of forming a phenolic resin type char upon thermal treatment.²³
14 Independent investigations of C3 upon heating has shown interesting results. Microscopic
15 analysis shows that sickle shaped microcrystals of C3 upon thermal treatment aggregate
16 (Figure 8b), fuse and undergo particle level intumescence (each particle swells). We pre-
17 assume that this process is also happening in the C3-TPU degradation but the process is
18 complicated because of the presence of copper ions probably in different oxidation state and
19 also because of the TPU matrix. There is also evidence (from py-GC/MS) of the presence of
20 species like 4-phenoxyphenol in the volatiles which is an intermediate or a degradation product
21 of poly-condensation of phenolic type resin which is also seen in hydroxyl *Salen*-TPU
22 degradation.^{23, 34} Thus, char formation has contribution from structural transformation of C3
23 and stable fragments from unzipped TPU matrix mediated by C3.
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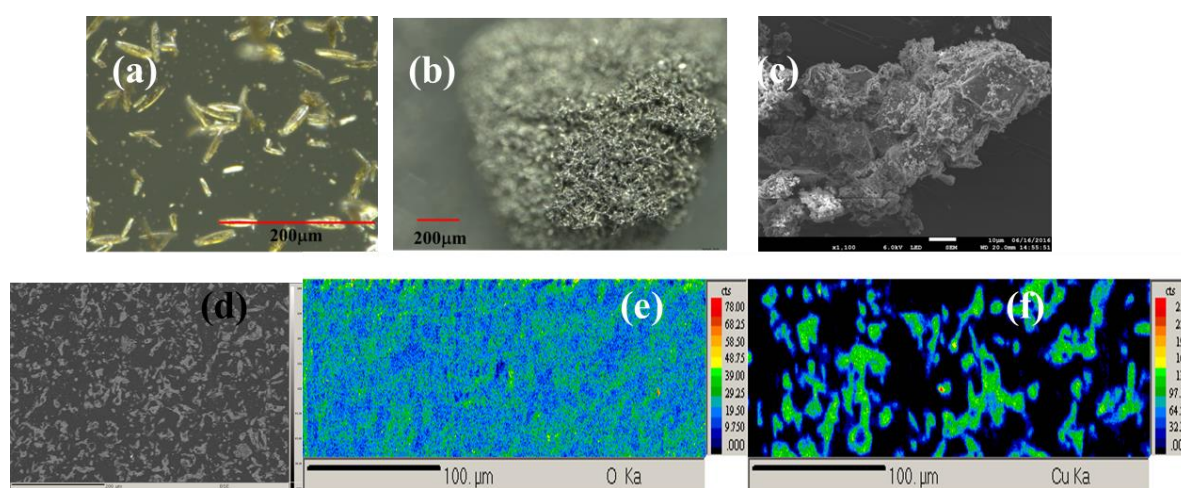


Figure 8. (a-c) Texture changes associated with C3 decomposition. (d-f) elemental mapping of the residue in C3 decomposition.

3.4.3 Fire behaviour investigation using EPR

3.4.3.1. Degradation study of C3-TPU under MLC experiment by CW EPR

In order to explain the structural changes of the complexes embedded in TPU and the radical formation, CW-EPR was carried out on C3-TPU and TPU samples. Figure 9 displays the CW-EPR spectra of powder collected at different MLC states i.e. onset, peak and end (residue) for C3-TPU and for neat TPU residue alone.

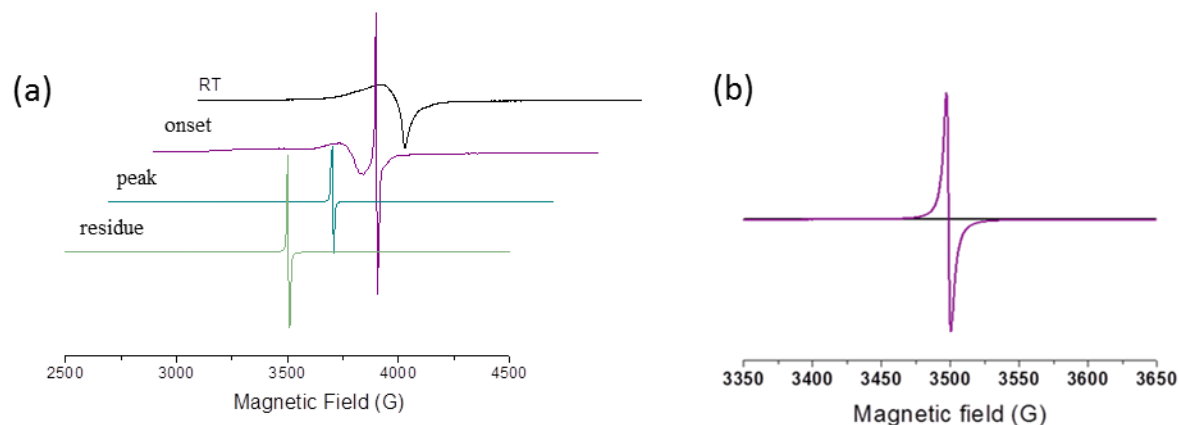


Figure 9. (a) CW-EPR spectra on C3-TPU degradation steps (sample collected at onset, peak and residues from MLC experiment). (b) CW-EPR spectrum of neat TPU (black line) and on residue (violet line).

As previously discussed, at room temperature C3-TPU spectrum shows a typical Cu (II) state signal. The onset spectrum permits to obtain the values of the two axial components of g factor ($g_{\perp} = 2.06$, $g_{\parallel} = 2.19$) together with the hyperfine constant $A_{\parallel} = 204$ G. As reported by Peisach and Blumberg⁴² the coordination of copper (II) is related to g_{\parallel} and A_{\parallel} values. These values reveal a copper with two nitrogen and two oxygen atoms coordination. In further MLC degradation steps, the Cu(II) signal disappears with a simultaneous formation of a strong and fine signal at 'g' value of 2.002, typical of organic radicals. Two hypotheses can explain the existence of radical species: the degradation of TPU matrix or the hydroquinone component of *Salen* framework which has been released from copper ions. In order to investigate the first suggestion, powders of neat TPU (without additives) were investigated by CW-EPR, before and after thermal treatment. EPR result presents in Figure 9b confirm the production of radical (g value of 2.002) at 600°C with the same EPR parameters observed before for C3-TPU. This value is characteristic of carbon centred radical. Radical formation has been reported in the degradation study of TPU.³⁷

At this stage, paramagnetic copper (II) ions are assumed to undergo a redox chemistry leading to change oxidation state of copper (0, I, III). Indeed, copper (I) and (III) are EPR silent species which explain the disappearance of copper signal in Figure 9a. This is consistent with the reported³⁶ work that showed the ability of the hydroxy-Salens to form the oxidizing Cu (III) species with the production of oxygen-based free radicals. XRD analysis (Figure 10) on the residue shows that indeed copper exists in its oxide form (Cu₂O and CuO peaks are found) and weak signal of metallic copper is also detected. During this degradation step, *Salen* framework detaches from copper ion and undergoes polycondensation leading resin type material. Experiments on C2-TPU showed similar conclusions like C3-TPU with the formation of radical species and the disappearance of copper in oxidation state (+II) (Figure S8).

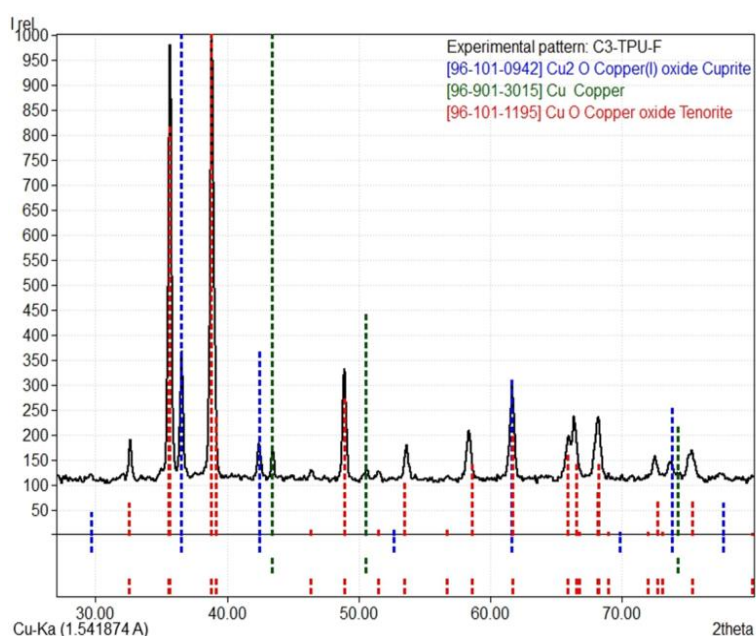


Figure 10. XRD of MLC C3-TPU residue.

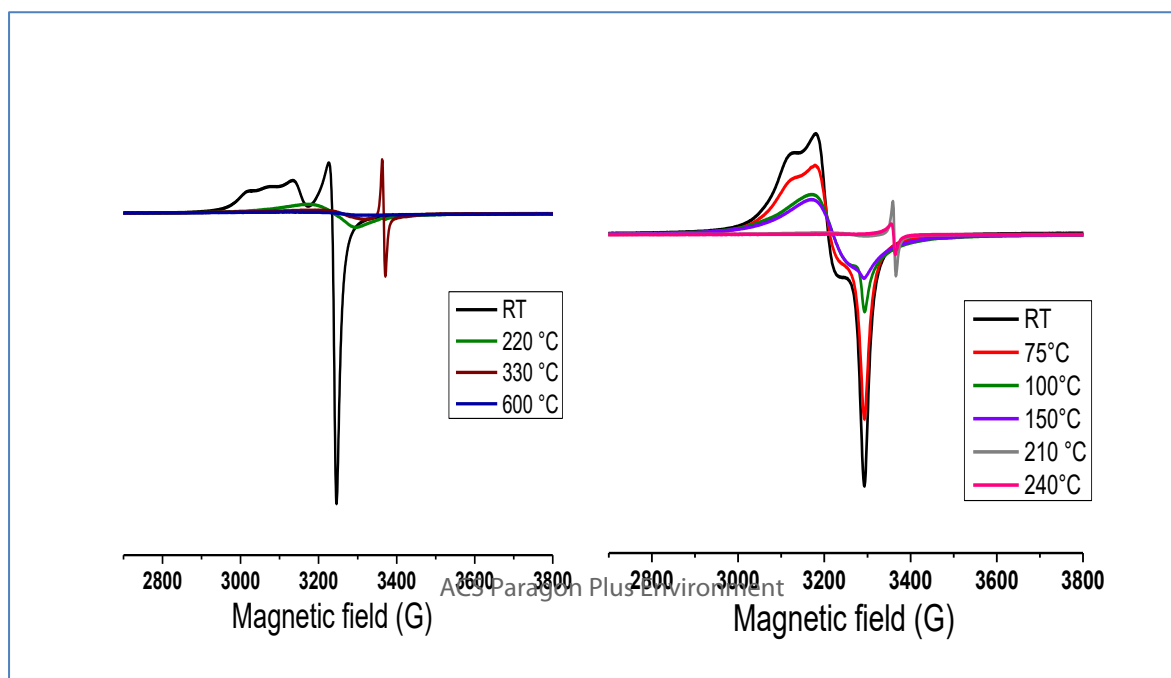
3.4.3.2. Investigation of the radicals origin: Characterization by *in situ* and *ex situ* CW EPR

The CW-EPR spectrum of C3-TPU final residue permits to support that the oxidation state of Cu(II) changes to Cu₂O, CuO and a less extend to Cu⁰ accompanied with the organic radical formation.

The second proposition concerning the hydroquinone component of *Salen* framework was to study the CW-EPR of C3 complex without TPU. To explain the degradation product of this complex, two types of EPR experiments were compared: *ex-situ* EPR experiments which correspond to the analyze of the residue collected from TGA experiment and *in-situ* EPR

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3 experiments to follow directly radical species formation during the heating of the sample inside
4 the EPR cavity. The EPR *in situ* approach permit to understand the origin of radical formation,
5 their evolution and their nature as a function of the temperature and will be correlated with the
6 *ex-situ* analyses. *In-situ* EPR characterization permit to obtain directly the relative quantity of
7 species formed all along the experiment. Figure 11 presents EPR spectra evolution during
8 thermal treatment. Changes in the EPR signals are similar on the both analysis. The initial EPR
9 signal of copper (II) decreases whereas the organic radical EPR signal increases. The difference
10 between signals shapes (*ex-situ* and *in-situ*) at room temperature is due to partial orientation of
11 powder in the EPR tube and caused by the large grain size and the small amount of powder. At
12 330°C, EPR spectrum (*ex-situ* mode) shows the presence of copper (II) with an axial symmetry
13 ($g_{\parallel}=2.19$ and $g_{\perp}=2.04$) and the hyperfine constant $A_{\parallel}=205$ G. In the meantime, the organic
14 radical with a fine and strong EPR signal ($g=2.004$) appears from 220 °C to 330°C and
15 disappears at 600°C. A radical recombination is suspected at higher temperature. These
16 experiments confirm an intermediate copper complex with two nitrogen and two oxygen atoms
17 coordination and the ability of copper complexes (without TPU) to produce radicals (probably
18 oxygen centered) during thermal treatment. The degradation of C3 alone during heating forms
19 radical species that are not able to maintain whereas radicals produced during the heating
20 degradation of C3-TPU are more durable (more stable). It is noteworthy that *Ex-situ* and *in-*
21 *situ* results give similar conclusions.

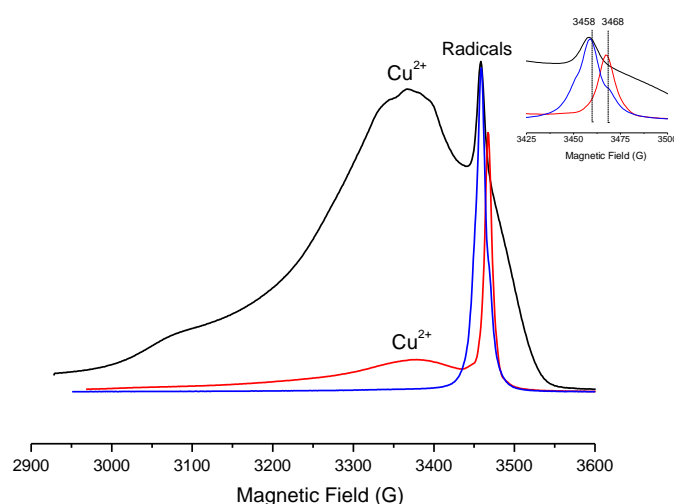
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These results suggest that the degradation of TPU at higher temperature contribute to create
stable radical species (carbon centered radicals) quite different from radicals from copper
complex (probably oxygen centered radicals).



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12 **Figure 11.** CW-EPR spectra: comparison between *ex-situ* (left) and *in-situ* (right) thermal
13 treatment of the neat C3 sample.
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16 3.4.3.3. Characterization of surrounding of the radicals by pulsed EPR

17 To go further in understanding of C3-TPU behaviour, additional experiments were performed
18 by using pulsed EPR spectroscopy on collected MLC samples (onset, peak and residue). 2-
19 pulses echo field sweep allows an orientation selection of the nuclear environments as observed
20 in Figure 12. Two paramagnetic species are present: Cu (II) is observed for C3-TPU-onset and
21 C3-TPU-peak while radicals are detected in all samples. Two types of radical species are
22 detected in the residues: the first one at 3458G (C3-TPU-onset, $g=2.0042$) and the second one
23 at 3468G (C3-TPU-peak, $g=2.0038$) while both species are identified for C3-TPU-residue
24 (radical at 3458G is predominant). As already seen by CW-EPR, C3-TPU-residue possesses a
25 spectral fingerprint characteristic of organic radicals. Copper (II) is not observed which
26 confirms a change of the oxidation state during the heat treatment. PEANUT experiment on
27 C3-TPU-residue - wherein both radicals are detected - helps to reveal electronic structures of
28 these radicals that give a direct access to spin state. The nutation spectrum is dominated by two
29 intense peaks at $\omega_{\text{TN}}^{(A)}/2\pi = 7.2$ MHz and $\omega_{\text{TN}}^{(B)}/2\pi = 9.9$ MHz (Fig. 13).
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54 **Figure 12.** 2-pulses echo field sweep recorded at 5 K on the C3-TPU-onset (black line), C3-
55 TPU-peak (red line) and C3-TPU-residue (blue line) samples.
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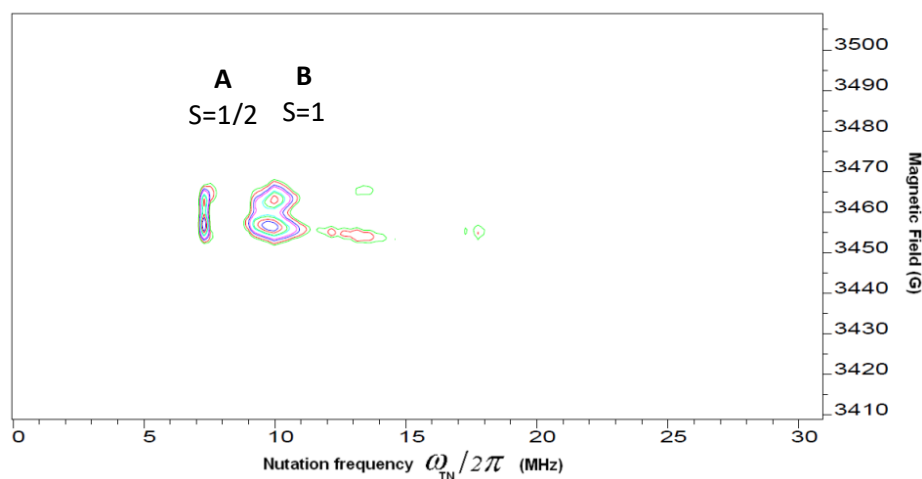


Figure 13: Transient nutation (PEANUT) experiment recorded on the C3-TPU-residue sample at 5 K.

We expect $\omega_{\text{TN}} = g\beta B_1/h$ for radicals with $S = 1/2$ (single electron) and $\omega_{\text{TN}} = \sqrt{2}g\beta B_1/h$ for biradicals with $S = 1$ (two unpaired electrons). For the same type of radical (with the same g -factor) and the same incident microwave field B_1 , the nutation frequency of a biradical ($S = 1$) is $\sqrt{2}$ times larger than for a single radical ($S = 1/2$). Nutation frequencies of (A) and (B) shown on Fig. 13 are related by $\omega_{\text{TN}}^{(\text{B})}/\omega_{\text{TN}}^{(\text{A})} = 9.9/7.2 = 1.375 \cong \sqrt{2}$, which corresponds to $S = 1$ and $S = 1/2$ for (B) and (A), respectively. Also, it may be concluded that the existence of two nutation peaks with frequencies in a $\sqrt{2}$ ratio is a strong argument for the attribution of spins $S = 1/2$ and 1 to (A) and (B) species, respectively. This technique allowed demonstrating the existence of two different organic radicals in two spin state: one with single electron and another one with two unpaired electrons which correspond to biradical (two different organic radicals with $\Delta g = 0.0004$). By combining the 2-pulses echo field sweep with PEANUT experiments, the results seem to show the presence of two radicals which are formed during the thermal treatment of C3-TPU. The first one is probably a carbon centered radical ($S = 1/2$) that is coming from the degradation of TPU which is more stable (mostly present in onset and residue samples) while the second one, probably an oxygen centered radical (biradical $S = 1$) is from the degradation of complex alone (mostly present on peak sample-intermediate temperature).

Further it is useful to investigate how local nuclear environments of copper and radicals undergo changes at different stages of degradation. Thus, two-dimensional HYSORE experiments are carried out. Spectra characteristic of the evolution of local surrounding of copper and organic radicals in all samples are presented in Fig. 14 and 15, respectively.

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3 The evolution of local chemical surrounding of Cu (II) in different samples collected from
4 MLC experiment is shown in Fig.14. Whatever the sample, a strong coupling with ^{14}N are
5 observed in (-,+) quadrant with coupling constant A of 6.6 MHz. This means that the ^{14}N nuclei
6 are very close to the unpaired electron of copper. In the (+,+) quadrant a pair of cross peaks
7 are centered at ^{14}N , ^{13}C , ^1H and ^{63}Cu nuclear Larmor frequencies are detected. All these nuclei are
8 weakly coupled with the unpaired electron of copper (onset and peak samples). A drastic
9 change in chemical environment with a strong contribution of nitrogen atoms at higher
10 temperature was observed. Indeed, the ratio C/N is very different for both samples: C3-TPU-
11 onset and C3-TPU-peak. C3-TPU-onset presents two types of ^{14}N nuclei characterized by
12 coupling constants A of 6.6 (strong coupling) and 1.4 MHz (weak coupling) while C3-TPU-
13 peak possesses coupling constants A of 6.6 (strong coupling) and 5 MHz (weak coupling) with
14 the presence of NH group. The flat ridges of protons observed arises mainly from Fermi contact
15 term corresponding to aliphatic protons. Cu (II) is coupled with 3 types of protons atoms
16 characterized by couplings constant A of 3, 7 and 20 MHz corresponding to different
17 hydrogens. It seems also that Cu (II) is coupled with diamagnetic copper atom which proves
18 the rapid change of the oxidation state of a part of copper during the degradation of onset
19 sample. For this sample, it is probable that a part of copper is oxidized into Cu (III) which is in
20 proximity of Cu (II).

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22 Figure 15 presents the evolution of local chemical environment of organic radicals. The carbon
23 contribution is very significant in the chemical surrounding of radicals (coupling constants of
24 1MHz for C3-TPU-onset and 5MHz for C3-TPU-peak and residue) which means that the local
25 microenvironment of radical species is carbonaceous. Despite the small natural abundance of
26 ^{13}C (1.11%), an intense signal of carbon is observed in fig.14. This is due to the large number
27 of carbon atoms located at short distances from the radical species and so, increases the
28 probability for the unpaired electron to interact with a ^{13}C nucleus. Consequently, the
29 probability of detecting the ^{13}C signal becomes high as the number of equivalent carbons
30 increases. It seems that the organic radicals are localized on large molecule as char. C3-TPU-
31 onset and C3-TPU-peak are composed of two types of nitrogen atoms (A=6.6 and 1.4 MHz)
32 while C3-TPU-residue exhibits a third type of nitrogen atom characteristic of NH group.
33 Different types of protons are also observed with similar coupling constants (A=3, 7MHz and
34 20MHz) but with an increase of dipolar part which corresponds to aromatic protons. Indeed,
35 the thermal treatment leads to an aromatization of the carbon chains.

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3 Thus, it was observed a drastic change in chemical environment with a strong contribution of
4 nitrogen atoms in sample collected at higher degradation stage from MLC which means that
5 the residues contain nitrogen fragments of TPU or Salen framework bearing nitrogen atoms. In
6 our earlier work²³ we have seen that nitrogen was found to be present in the residue of *Salen*
7 degraded product and we proposed that benzoxazine type resin formation from *Salen*
8 molecules. Thus, the char are nitrogen rich residues.
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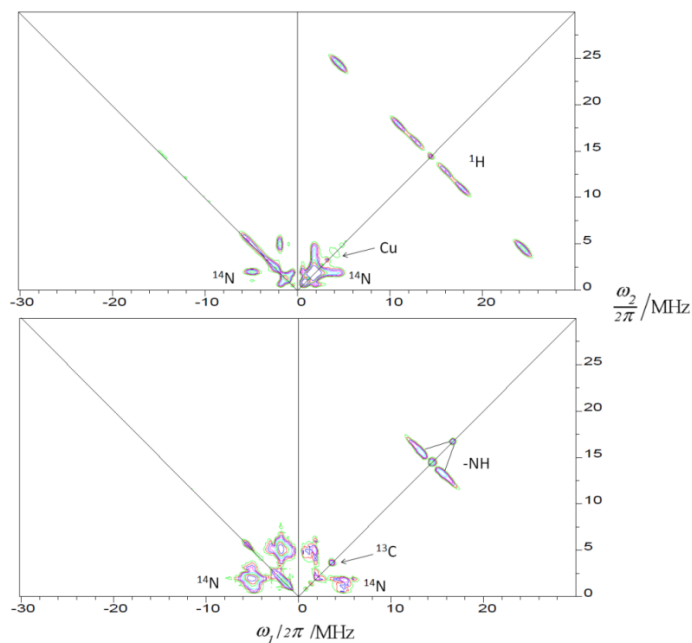


Figure 14. Evolution of the local chemical surrounding of Cu (II) by 2D-HYSCORE recorded on the C3-TPU-onset (top) and C3-TPU-peak (bottom) at 5 K and 3367 G.

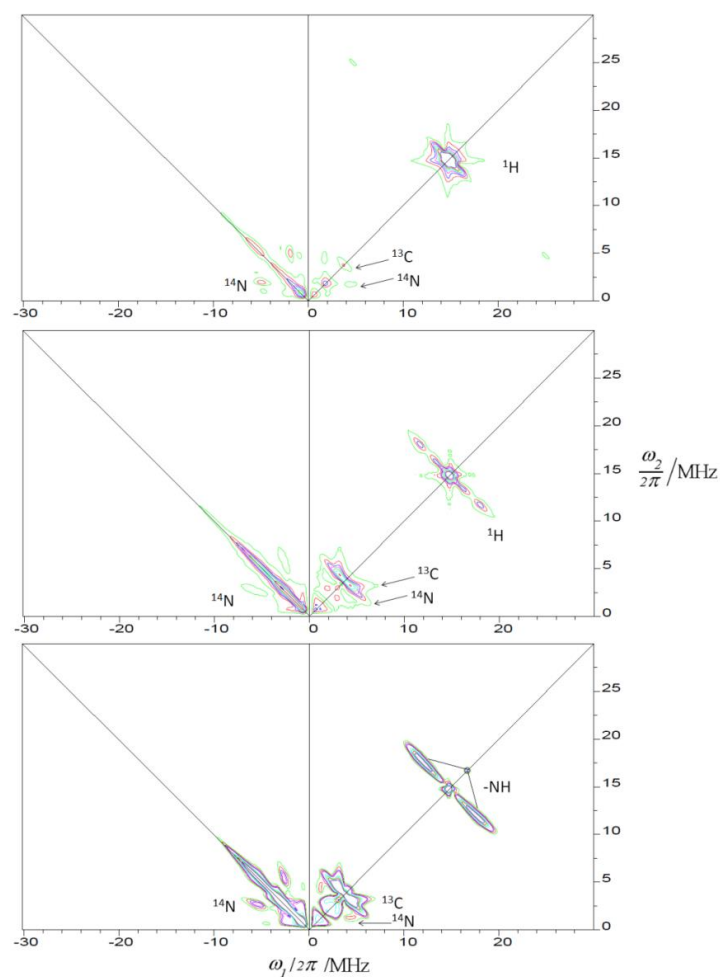


Figure 15. Evolution of the local chemical surrounding of organic radicals by 2D-HYSCORE recorded on the C3-TPU-onset (top), C3-TPU-peak (middle) and C3-TPU-residue (bottom) at 5 K and 3458 or 3468 G.

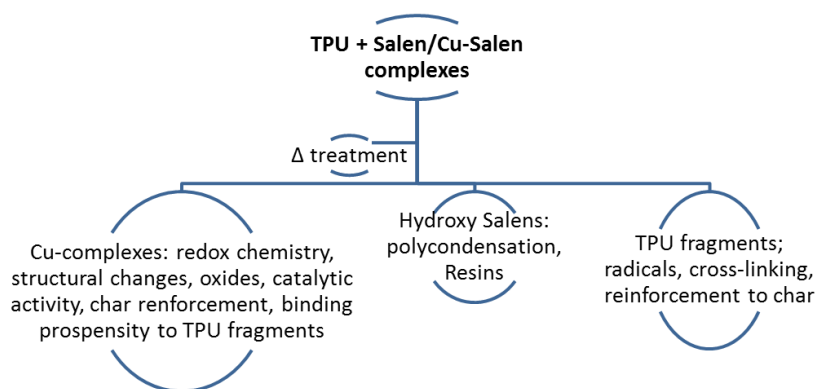
3.4.3.4. Role of copper ions in the char

Following section is dedicated to additional role of copper ions in the char. From the ongoing section, we also understood that copper oxides reinforce char. As copper oxide presence is prominent in the char and considering the reported^{10,11,13,14,41} catalytic role of oxides of copper we have studied the carbon monoxide production profile for TPU and TPU formulation. Results are obtained from MLC-FTIR experiment and figure S11 shows CO production profile for neat TPU, salen based TPU formulation (L2-TPU) and C2-TPU. It shows that TPU's CO evolution is shared over 3 regions. The middle being the highest (~600 Vppm) and the last one is very broad (350 Vppm) and extends to the end. The curve is similar to TPU at the beginning in L2-TPU but drastically changes at later stages with significant increase in CO production. However, in C2-TPU the profile is very different (Figure S11). There is only one major peak at the beginning, like for TPU and L2-TPU, continuously decreasing as the decomposition

progresses. The decreasing is assumed to be due to the catalytic activity of cuprous oxide (different copper species are identified in XRD) in the conversion of carbon monoxide.

4. Mode of action:

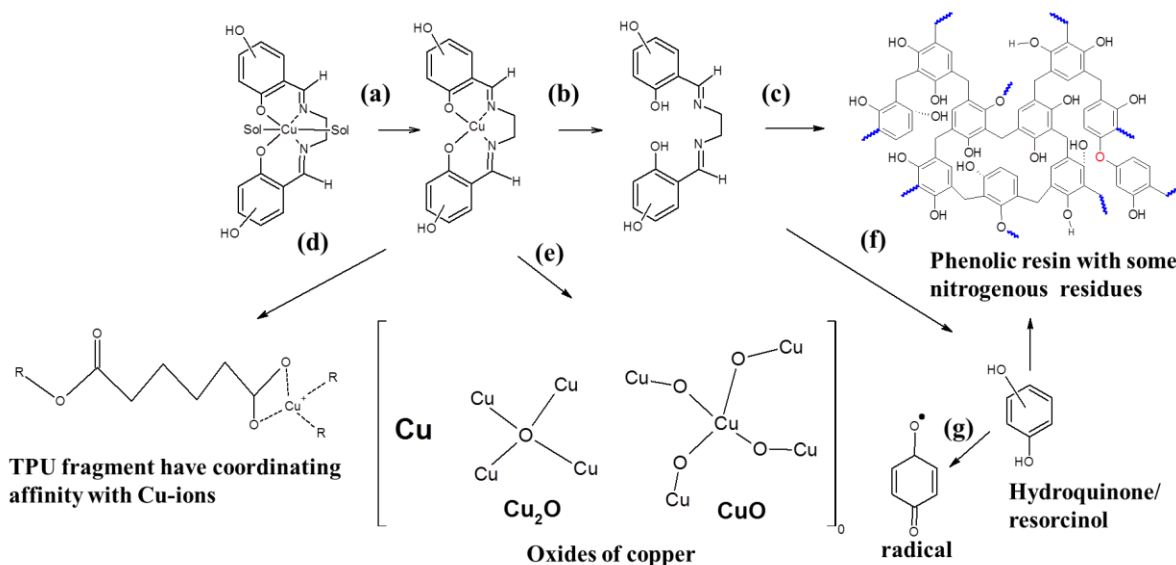
Based on the evidences obtained so far, the intumescence formation with C2-TPU or C3-TPU is found to be a combined contribution of copper complexes and TPU fragments that have cross-linking tendencies (Scheme 1). In our earlier work we have shown that the ligand of copper complexes play important role in the fire protection mechanism.²³ Advantage of copper complexes here is the better thermal stability over its metal free ligand (by 30-40°C) and involvement of its redox chemistry which are discussed in the following section.



Scheme 1. An overview of participation of components of formulation in fire protection mechanism.

Hydrogen bonding connecting additives and TPU are loosen as the TPU matrix with additives is subjected to thermal treatment. Copper *Salen* starts to release coordinated water molecules. Further copper centres undergo redox reaction with subsequent change in oxidation states, coordination number and geometry before releasing into the matrix. This will be simultaneously accompanied by release of free hydroxyl *Salen* framework which is a potential molecule in the formation of heat resistant resins. Thereafter hydroxyl *Salen* cannot stay in its neat form but can undergo fragmentation as observed before to form reactive hydroxyl substituted phenols or *o*-quinone methide (*o*-QM) like intermediates in the resin formation pathway.²³ The advantage of this fragmentation after demetalation is that it does not end up with decomposition of organic framework of ligand but instead steady organic transformation.

In the present case, poly-condensation leading to phenolic resin type thermal protecting barrier is evidenced (Scheme 2). Thus, copper complexes can contribute as a precursor of phenolic resins. The whole process will influence the TPU unzipping pathway. Hydroxyl *Salen* during its polycondensation process also form some amount of benzoxazine or nitrogen containing resins. Collectively these resins are the source of char. C2 and C3 cannot direct same type of resin formation pathway in their TPU matrix as observed in case of metal free 4- and 5-hydroxy *Salen* which followed different routes in resin formation pathway.^{22,23}

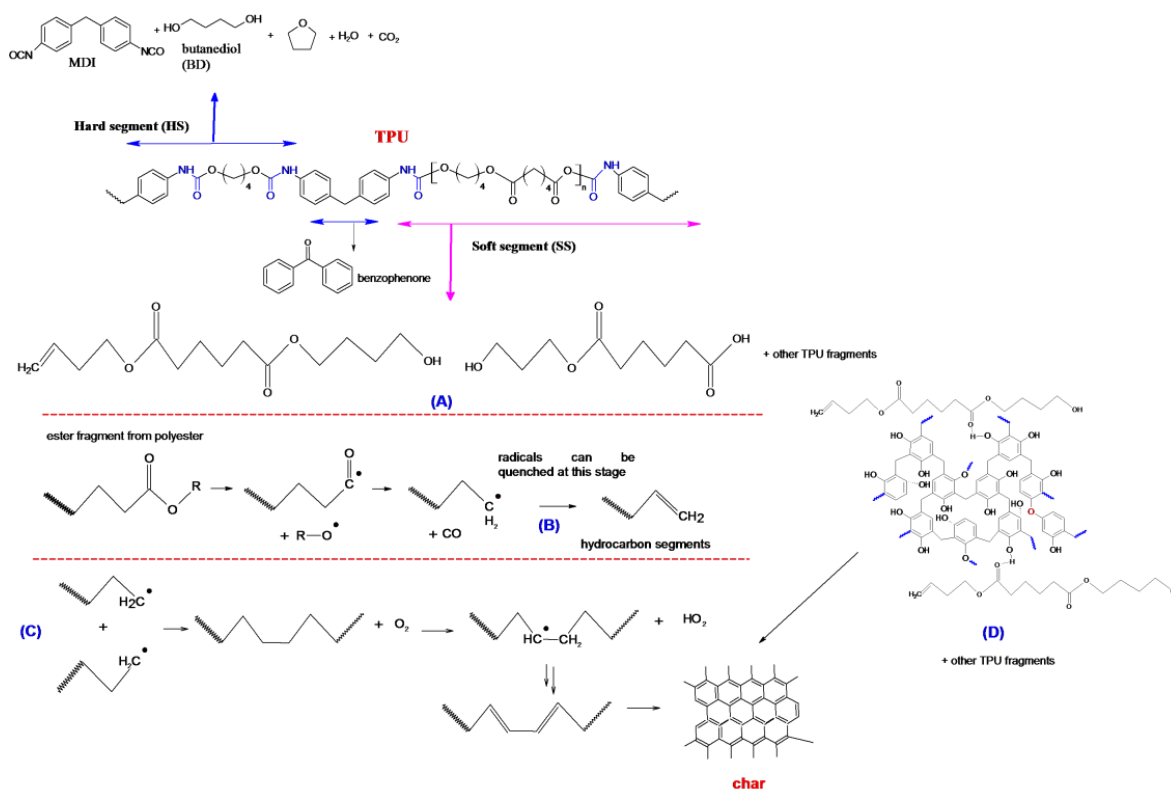


Scheme 2. Tentative degradation pathway of copper complexes in TPU.

Meanwhile, copper ions released into the matrix undergo several changes particularly oxide formation and integration into the char formation. Metal ions also can coordinate to fragmented section of TPU (with potential donor molecules) and keeps them in condensed phase (Scheme 2). Further advantage of formation of copper oxide lies in its catalytic activity. Cuprous oxide is known to exhibit synergistic effect on improving the flame retardancy of resins or polymers.⁴¹ Redox active species like Cu₂O enters the redox cycle and catalyse the conversion of CO into CO₂ and also acts as smoke suppressant which has been proved through both experimental and theoretical studies.⁴¹

TPU unzipping can follow more than one type of polymer unzipping. Primary target being the urethane group it is found that hard segment starts to break forming methylene diphenyl

diisocyanate (MDI) and butane diol (BD). Although the primary target of degradation is urethane group, there is report that urethane segment is unaffected and methylene group of MDI oxidation forms benzophenone.³⁷ Further soft segment of TPU having ester fragment undergoes thermo-oxidative degradation giving oligomers and some cyclized products depending upon the TPU unzipping process.^{37,38} MDI once released can act as cross-linker and thus retained in the condensed phase. TPU unzipping process under these circumstances may favour the production of larger TPU fragments and less combustible materials. These fragments having potential donor atoms encapsulate copper ions and contribute to the matrix of char.



Scheme 3. Tentative and simplified degradation pathway in formulations.

Radical based polymerisation of hydrocarbon residues from TPU is known and is also expected to contribute to the char formation in present case. The free radical cross-linking of TPU³⁹ fragments is considered to involve the generation of free radical positions in the polyurethane chains. Methylene sites are suggested for free-radical formation. Initially ester segment breaks between carbonyl-O bond and produce carbonyl radical and alkoxy radical evolving to methylene radical and the coupling of these radicals takes place (Scheme 3) to form C-C covalent chemical cross-links. Propagation of radical formation can be quenched by FR²⁴ or its fragments and thus formation of some smaller segment of hydrocarbons is possible which can

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3 lead to dripping. On the other hand, if free radical addition proceeds unopposed then the process
4 of cross linking of segments proceed with the formation of char (passing through stages such
5 as diene, polyene, polyaromatics and finally char) (scheme 3). These above two processes are
6 respectively happening in C3-TPU and C2-TPU. We have seen from EPR studies that radical
7 generation is possible both with hydroxyl *Salen* and also with TPU matrix. Additional EPR
8 results described on radicals are helpful particularly in understanding spin states of radicals (by
9 PEANUT experiment) and local nuclear environment changes (by 2D-HYSCORE
10 experiment).

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12 Although intumescence is evidenced during the fire test on C2-TPU but it is not strong enough
13 to sustain for longer time. It is due to surface quality of char and overall physical strength of
14 char particularly upper surface of char. It is observed that there are few red patches over the
15 intumescence (Figure 16 c). Sample after LOI test also shows this tendency of formation of red
16 patches. This area on char (e.g. char from MLC) cracks under high heat flux and are possibly
17 responsible for the second pHRR (Figure 2a). Residues containing this red patches were
18 analysed by XRD (Figure 10) which shows formation of CuO, Cu₂O and metallic copper.
19 Copper oxide is known to integrate into the char and is assumed to strengthen the char.
20 However precipitation of metallic copper (Cu) fails to integrate into chars into chars forming
21 porous regions created by Cu metal. This poor cohesion of intumescence leads to cracks and
22 pores (Figure 16). Also such regions on char pass through a stage of smouldering effect that
23 (prolonged localised heating) oxidises the char at localized points precipitating metallic copper.
24 This weakens the char.



50 **Figure 16.** (a) Char formation as seen in MLC experiment in C2-TPU. (b) Sample collected at
51 the onset of experiment (c) red patches on the char on the char collected at 'onset' (d,e)
52 Focussed images on red patches.
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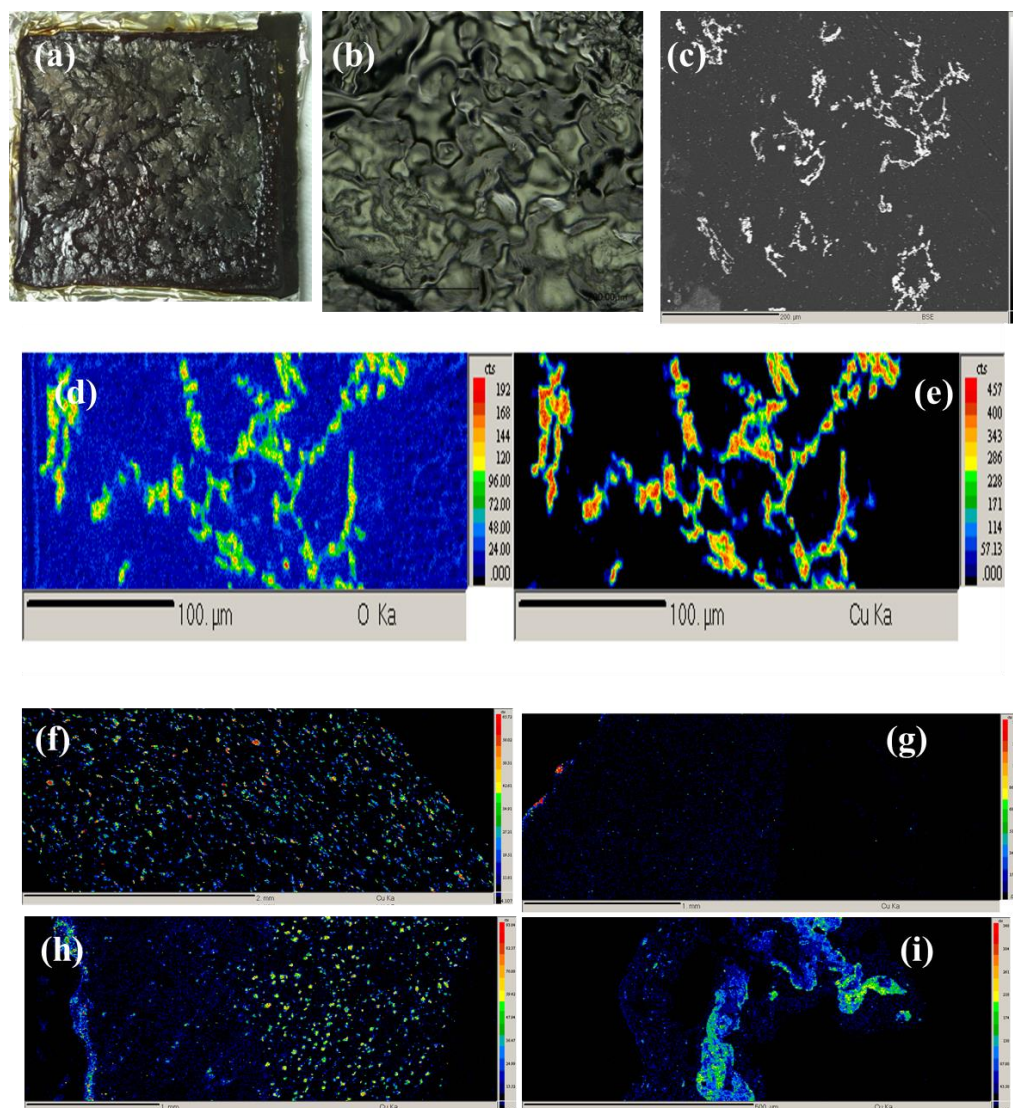


Figure 17. Textural changes in C3-TPU. (a) C3-TPU sample from cone calorimetry experiment stopped at peak of pHRR. (b) Focussed image (c) BSE image on residue from TGA (d,e) elemental mapping on residue (f-i) EPMA on residue showing localization of copper complexes in the sample. (f) Neat C3-TPU. (g) C3-TPU-onset. (h) C3-TPU-peak. (i) C3-TPU-char.

On the other hand the occasional red patches seen in C2-TPU over the intumescence which is due to precipitation of metallic copper are not seen in C3-TPU and their deleterious effect is minimized due to their submergence in highly viscous swelled /flexible char (Figure 16 a,b). It should be reminded that L3 also mainly works by a kind of dripping (highly viscous residue) and follow different type of phenolic resin formation.²⁴ This kind of flexible char under high heat flux does not cracks and possibly responsible for delayed TTI and pHRR reduction. Thus, C3-TPU works mainly by flexible char formation. Another interesting observation is the

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3 formation of copper depletion zone in the residues (residue from MLC experiment). It is due
4 to migration (Figure 17f) of additives to the surface upon exposure to high temperature which
5 then leaves behind resin forming hydroxyl Salen and concentrate at the surface by reinforcing
6 char surface.
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11 Although there is not significant contribution to gas phase action certain observations are worth
12 noticing. We have seen from PCFC that both peaks of C2-TPU are reduced compared to neat
13 TPU. This reduction is not seen in other complexes. Therefore, there is some contribution to
14 the gas phase action of flame retardancy. This could be passive pathway of reduction in
15 flammable material or active way of quenching of radicals and/or dilution of combustible
16 material. If we compare weight loss in the first step of TGA (16%), it is comparatively lower
17 in C2-TPU meaning promotion of char formation. From Py-GC/MS results, it is interesting to
18 note that cyclopentanone produced is greatly minimised indicates changes in the mode of
19 polymer unzipping. Resorcinol which is a known radical scavenger is clearly detected in py-
20 GC/MS of neat C2 although its detection in C2-TPU is difficult. This may also contribute to
21 flame retardancy either in the gas phase or condensed phase by radical scavenging.⁴⁰ Thus it
22 seems that the reduction in gas phase is partially due to reduction of flammable material.
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32 Thus far, we have seen that how copper complexes undergo structural changes upon thermal
33 treatment maximising char formation, minimizing formation of flammable products,
34 contributing in the oxide forms to the flame retardancy.
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39 4. Conclusion:

40 Fire protection by *Salen* complexes described in this work underline the advantages of using
41 metal complexes as fire retardants. Complexation not only increase the thermal stability
42 compared to un-complexed *Salen* but also provide a platform for structural transformation
43 forming resin like insulating materials. Metal ions released during heat treatment reinforce char
44 and have some catalytic role in the smoke production. Another advantage of complexation is
45 to avoid unnecessary coordination of metal ions (when used in uncomplexed form) with
46 potential donor sites on polymer thus this lower the risk of altering the physical and chemical
47 properties of the matrix. *Salen* in our case works without a synergist and mostly perform in the
48 condensed phase although gas phase action could not be ruled out. However, additional
49 synergist could even enhance the overall efficiency. Moreover, by the variation of the
50 functional groups on the FR, the secondary interaction and molecular organization with the
51 host matrix can be controlled and thus its functional properties can be influenced. Redox active
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3 metal ions (and also metal ions like lanthanides with higher coordination numbers) involving
4 paramagnetism, although not as easy as diamagnetic metal ions to follow, may have curious
5 but interesting properties during their change in oxidation states. We have seen in the present
6 study how EPR techniques could be an important tool as complementary technique of solid
7 state NMR in characterising the steps involved in elucidating the flame retardancy. Present
8 study has demonstrated that formulations with paramagnetic metal ions and radical formation
9 can be easily tracked by EPR. Further advanced EPR techniques like PEANUT, 2D HYSCORE
10 can be explored to study spin states of radicals and local environment of radical respectively in
11 the elucidation of mechanism of action of flame retardancy. Overall, the study has
12 demonstrated that tailored classical metal complexes will have huge prospects in flame
13 retardancy.
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24 CNRS 3443). We acknowledged IRCP laboratory which permitted us to perform High
25 Temperature EPR studies.
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32 References:

- 33 1. Morgan, A.B.; A Review Of Transition-Metal Based Flame Retardants: Transition Metal
34 Oxides/Salts, And Complexes, Fire And Polymers V, Chapter 19, pp 312-328, ACS
35 Symposium Series, Vol 1013.
- 36 2. Levchik, S.V.; Levchik, G.F.; Camino, G; Costa, L; Lesnikovich, A.I. Mechanism Of
37 Action Of Phosphorus-Based Flame Retardants In Nylon 6. III. Ammonium
38 Polyphosphate/Manganese Dioxide Fire Mater. 1996, 20,183-190.
- 39 3. Davies, P.J.; Horrocks, A.R.; Alderson, A. The Sensitization Of Thermal Decomposition
40 Of Ammonium Polyphosphate By Selected Metal Ions And Their Potential For Improved
41 Cotton Fabric Flame Retardancy. Polym. Degrad. Stab. 2005, 88, 114–122.
- 42 4. Wang, D.Y.; Liu, Y; Wang, Y.Z.; Artiles, CP.; Hull TR.; Price D. Fire Retardancy Of A
43 Reactively Extruded Intumescent Flame Retardant Polyethylene System Enhanced By
44 Metal Chelate. Polym. Degrad. Stab. 2007, 92, 1592–1598.
- 45 5. Wang, D.Y.; Liu, Y; Ge, X.G.; Wang, Y.Z.; Stec, A; Biswas, B; Hull; R.T.; Price, D.
46 Effect Of Metal Chelates On The Ignition And Early Flaming Behaviour Of Intumescent
47 Fire-Retarded Polyethylene Systems. Polym. Degrad. Stab. 2008, 93,1024–1030.
48
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 - 57
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 - 60
6. Tian, C.M.; Qu, H.Q.; Wu, W.H.; Guo, H.Z.; Xu, J.Z. Metal Chelates As Synergistic Flame Retardants For Flexible PVC J. Vinyl Addit. Technol. 2005, 11, 70–75.
 7. Nodera, A; Kanai, T. Thermal Decomposition Behavior And Flame Retardancy Of Polycarbonate Containing Organic Metal Salts: Effect Of Salt Composition J. Appl. Polym. Sci. 2004, 9, 2131–2139.
 8. Wang, D.L.; Liu, Y.; Wang, D.Y.; Zhao, C.X.; Mou, Y.R.; Wang, Y.Z. A Novel Intumescent Flame-Retardant System Containing Metal Chelates For Polyvinyl Alcohol Polym. Degrad. Stab. 2007, 92, 1555–1564.
 9. Wu, N.; Yang, R. Effects Of Metal Oxides On Intumescent Flame-Retardant Polypropylene, Polym. Adv. Tech. 2011, 22, 495–501.
 10. Zhou, Y.; Liu, X.; Wang, F.; Hao, J.-W.; Du, J.-X. Effect Of Metal Oxides On Fire Resistance And Char Formation Of Intumescent Flame Retardant Coating J. Inorg. Mater. 2014, 29, 972-978.
 11. Wang, L., Yang, W.; Wang, B.; Wu, Y.; Hu, Y.; Song L.; Yuen; R. K. K. The Impact Of Metal Oxides On The Combustion Behavior Of Ethylene–Vinyl Acetate Copolymers Containing An Intumescent Flame Retardant, Ind. Eng. Chem. Res., 2012, 51, 7884–7890.
 12. Wu, J.; Hu, Y.; Song, L.; Kang, W. Synergistic Effect Of Lanthanum Oxide On Intumescent Flame-Retardant Polypropylene-Based Formulations; J. Fire Sci. 2008, 26, 399-414.
 13. Metal Catalysed Intumescence Of Polyhydroxyl Compounds, Phd Thesis, 2003, University Of Pretoria, By F.J.W.J. Labuschagne.
 14. Lin, M.; Li, B.; Li, Q.; Li, S.; Zhang, S. Synergistic Effect Of Metal Oxides On The Flame Retardancy And Thermal Degradation Of Novel Intumescent Flame-Retardant Thermoplastic Polyurethanes, Applied Polymer Sci. 2011, 121, 1951–1960.
 15. Fontaine, G.; Turf, T.; Bourbigot, S. Salen Copper Complexes: A Novel Flame Retardant For Thermoplastic Polyurethane, Fire And Polymers V, Chapter 20, Acs Symposium, 2009, 1013, 329.
 16. Sheng, W. J.; Liu, Y.; Zhao, H.-B.; Liu, J.; Wang, D.-Y.; Song, Y.-P.; Wang, Y.-Z. Metal Compound-Enhanced Flame Retardancy Of Intumescent Epoxy Resins, Containing Ammonium Polyphosphate, Polym. Degrad. Stab. 2009, 94, 625–631.
 17. Naik, A.D.; Fontaine, G.; Samyn, F.; Delva, X.; Bourgeois, Y.; Bourbigot, S. Melamine Integrated Metal Phosphates As Non-Halogenated Flame Retardants: Synergism With

- 1
2
3 Aluminium Phosphinate For Flame Retardancy In Glass Fiber Reinforced Polyamide 66
4 Polym. Degrad. Stab. 2013, 98, 2653-2662.
5
6
7 18. Naik, A.D.; Fontaine, G.; Samyn, F.; Delva, X.; Louisy, J.; Bellayer, S.; Bourgeois, Y.;
8 Bourbigot, S. Outlining The Mechanism Of Flame Retardancy In Polyamide 66 Blended
9 With Melamine-Poly(Zinc Phosphate) Fire Safety J. 2014, 70, 46-60.
10
11
12 19. Gardelle, B.; Duquesne, S.; Vandereecken, P.; Bellayer, S.; Bourbigot, S. Resistance To
13 Fire Of Intumescent Silicone Based Coating: The Role Of Organoclay Prog. Org. Coat.,
14 2013, 76, 1633–1641.
15
16
17 20. Kilinc, M. Silicon Based Flame Retardants. In: Morgan A.B. And Wilkie C.A. (Eds)
18 Non-Halogenated Flame Retardant Hand Book. Hoboken, NJ: John Wiley & Sons, Inc.,
19 2014.
20
21
22 21. Fontaine G.: Metal-Organic Frameworks : New Fire Retardant For Polymeric Materials,
23 F.R.P.M. 2017, Manchester, Uk
24
25
26 22. Naik, A. D.; Fontaine, G.; Bellayer, S.; Bourbigot, S. Salen Based Schiff Bases To Flame
27 Retard Thermoplastic Polyurethane Mimicking Operational Strategies Of Thermosetting
28 Resin, RSC Adv. 2015, 5, 48224-48235.
29
30
31 23. Naik, A. D.; Fontaine, G.; Bellayer, S.; Bourbigot, S. Crossing The Traditional
32 Boundaries: Salen-Based Schiff Bases For Thermal Protective Applications, ACS Appl.
33 Mater. Interfaces 2015, 7, 21208-21217.
34
35
36 24. Chattopadhyay, D.K.; Webster, D.C. Thermal Stability And Flame Retardancy Of
37 Polyurethanes Prog. Polym. Sci. 2009, 34, 1068-1133.
38
39
40 25. Tabuani, D.; Bellucci, F.; Terenzi, A.; Gamino, G. Flame Retarded Thermoplastic
41 Polyurethane (TPU) For Cable Jacketing Application Polym. Degrad. Stab., 2012, 97,
42 2594.
43
44
45 26. HöFer, P.; Grupp, A.; Nebenführ, H.; Mehring, M. Hyperfine Sublevel Correlation
46 (Hyscore) Spectroscopy: A 2d ESR Investigation Of The Squaric Acid Radicalchem
47 Phys. Lett. 1986, 132, 279–282.
48
49
50 27. Astashkin, A. V.; Schweiger A. Electron-Spin Transient Nutation: A New Approach To
51 Simplify The Interpretation Of ESR Spectra. Chem. Phys. Lett. 1990, 174, 595–602.
52
53
54 28. Baker, E.N.; Hall, D.; Waters, T.N. The Color Isomerism And Structure Of Copper Co-
55 ordination Compounds. Part X. The Crystal Structure Of Bis-Salicylaldiminatocopper
56 (II) J. Chem. Soc., A, 1966, 680-684.
57
58
59 29. <http://www2.basf.us/urethanechemicals/tpu/pdf/c85a.pdf>
60

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 - 46
 - 47
 - 48
 - 49
 - 50
 - 51
 - 52
 - 53
 - 54
 - 55
 - 56
 - 57
 - 58
 - 59
 - 60
30. Rosu, D.; Tudorachi, N.; Rosu, L. Investigations On The Thermal Stability Of A Mdi Based Polyurethane Elastomer Anal. J. Appl. Pyro. 2010, 89, 152.
31. Moo Espinosa, J.I.; Cauich-Rodríguez, J.V.; Ávila-Ortega, A.; Vázquez-Torres, H.; Marcos-Fernández, A.; San Román, J. Cervantes-Uc J.M. TGA/FTIR Studies Of Segmented Aliphatic Polyurethanes And Their Nanocomposites Prepared With Commercial Montmorillonites, Polym. Degrad. Stab. 2009, 94, 1666-1677.
32. Herrera, M.; Matuschek, G.; Kettrup, A. Thermal Degradation Of Thermoplastic Polyurethane Elastomers (TPU) Based On MDI Polym. Degrad. Stab. 2002, 78, 323-331.
33. Berta, M.; Lindsay, C.; Pans, G.; Camino, G. Effect Of Chemical Structure On Combustion And Thermal Behaviour Of Polyurethane Elastomer Layered Silicate Nanocomposites Polym. Degrad. Stab. 2006, 91, 1179-1191.
34. Hesse, W.; Lang, J. Phenolic Resins, Ullmann's Encyclopedia Of Industrial Chemistry, 2000, 26, 583.
35. Schartel, B.; And Hull, T. R. Development Of Fire-Retarded Materials—Interpretation Of Cone Calorimeter Data. Fire Mater. 2007, 31, 327-354.
36. K.-H., McBride M.B. Electron Transfer Processes Between Hydroquinone And Iron Oxides Kung Clays Clay Miner. 1988, 36, 303-309.
37. Zaikov, G.E.; Lomakin, S.M. New Aspects Of Ecologically Friendly Polymer Flame Retardant Systems Polym. Degrad. Stab. 1996, 54, 223-233.
38. Servay, T.; Voelkel, R.; Schmiedberger, H.; Lehmann, S. Thermal Oxidation Of The Methylene Diphenylene Unit In MDI-TPU Polymer 2000, 41, 5247-56.
39. Handbook Of Elastomers, Second Edition, Bhowmick A.K.; Stephens H.; CRC Press, 2000, 265.
40. Thavasi, V.; Phillip, R. Bettens A.; Leong L. P. Temperature And Solvent Effects On Radical Scavenging Ability Of Phenols Phys. Chem. A 2009, 113, 3068-3077.
41. Chen, M-J.; Lin, Y-C.; Wang, X-N.; Zhong, L.; Li, Q-L; Liu, Z-G. Influence Of Cuprous Oxide On Enhancing The Flame Retardancy And Smoke Suppression Of Epoxy Resins Containing Microencapsulated Ammonium Polyphosphate, Ind. Eng. Chem. Res. 2015, 54, 12705-12713.
42. Peisach, J.; Blumberg, W.E. Structural Implications Derived From The Analysis Of Electron Paramagnetic Resonance Spectra Of Natural And Artificial Copper Proteins, Arch. Biochem. Biophys. 1974, 165, 691-708.

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Graphical Abstract

