

Revealing Brønsted Acidic Bridging SiOHAl Groups on Amorphous Silica-Alumina by Ultra-high Field Solid-state NMR

Zichun Wang,^{a,b} Kuizhi Chen,^c Yijiao Jiang,^b Julien Trébosc,^{d,e} Wenjie Yang,^a Jean-Paul Amoureux,^{d,f,g} Ivan Hung,^c Zhehong Gan,^c Alfons Baiker,^h Olivier Lafon,^{*,d,i} Jun Huang^{*,a}

^a Laboratory for Catalysis Engineering, School of Chemical and Biomolecular Engineering, The University of Sydney, NSW 2006, Australia.

^b Department of Engineering, Macquarie University, Sydney, New South Wales 2109, Australia.

^c National High Magnetic Field Laboratory, Tallahassee, Florida 32310, United States

^d Univ. Lille, CNRS, Centrale Lille, Univ. Artois, UMR 8181-UCCS-Unité de Catalyse et Chimie de Solide, F-59000 Lille, France.

^e Univ. Lille, CNRS, Centrale Lille, Univ. Artois, FR 2638, Federation Chevreul, F-59000 Lille, France.

^f Bruker Biospin, 34, rue de l'industrie, 67166 Wissembourg, France.

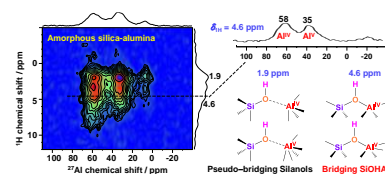
^g Riken NMR Science and Development Division, Yokohama, 230-0045 Kanagawa, Japan.

^h Institute for Chemical and Bioengineering, Department of Chemistry and Applied Bioscience, ETH Zürich, HCI, CH-8093 Zürich, Switzerland.

ⁱ Institut Universitaire de France.

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ABSTRACT: Amorphous silica-aluminas (ASAs) are important acidic catalysts and supports for many industrially essential and sustainable processes. The identification of surface acid sites with their local structures on ASAs is of critical importance for tuning their catalytic properties, but still remains a great challenge and under debate. Here, ultra-high magnetic field (35.2 T) ²⁷Al-¹H D-HMQC (dipolar-mediated hetero-nuclear multiple-quantum correlation) two-dimensional NMR experiments demonstrate two types of Brønsted acid sites in ASA catalysts. In addition to the known pseudo-bridging silanol acid site, the use of ultra-high field NMR provides the first direct experimental evidence for the existence of bridging silanol (BS: SiOHAl) acid sites in ASAs, which has been hotly debated in the past few decades. This discovery provides new opportunities for scientists and engineers to develop and apply ASAs in various reaction processes, due to the significance of BS in chemical and fuel productions based on its strong Brønsted acidity.



To realize green and sustainable energy and chemical productions, eco-friendly and cost-effective silica-alumina catalysts are extensively utilized for driving versatile acid-catalyzed hydrocarbon conversion reactions in drug and fine chemistry, petrochemical, and bio-refining industries.¹⁻² Compared to microporous zeolites, amorphous silica-aluminas (ASAs) offer unconstrained mass transfers and thus exhibit excellent performances in reactions involving large molecules as both acid catalysts and supports for metal active species. These properties render ASAs particularly interesting for applications in the conversion of biomass and derivatives for renewable energy and chemicals.³⁻⁸ Moreover, ASAs are widely utilized in large-scale hydrocracking processes to balance the acidity for the

selective production of middle distillates, such as diesel and kerosene fractions.⁹

So far, several models for the formation of catalytically active Brønsted acid sites (BASs) in ASAs have been proposed.^{3, 10-12} Two models are commonly considered for the formation of BASs on ASAs: 1) pseudo-bridging silanols (PBS: SiOH•••Al)^{3, 13-15} with silanol groups tilting towards Al atoms with weak interaction; and 2) silanols weakly bounded to neighboring Al sites via electrostatic interactions.¹⁶⁻¹⁷ In these models, there is no covalent bond between silanols and aluminium atoms. The acidic strength of these protons depends on Al sites coordinating to the hydroxyl oxygen of neighboring silanols as well as the SiOH•••Al distance. Due to their amorphous nature, ASAs

exhibit average SiOH•••Al distances in the range of 2.94–4.43 Å, which are significantly longer than for bridging OH groups (SiOH-Al) in crystalline zeolites (1.88–2.0 Å).¹³⁻¹⁴ These elongated SiOH•••Al distances could weaken the polarization of the hydroxyl oxygens by neighboring Al sites, accounting for the much lower Brønsted acidity and catalytic performance of ASAs compared to zeolites.¹⁸⁻¹⁹ In contrast, the weak acidity and poor activity of ASAs have been ascribed to the small amount of strong BASs existing in ASAs compared to zeolites.²⁰ However, the existence of strong BASs in ASAs has been reported by using reactivity measurements in propane cracking²¹ and H/D exchange reactions.²² Zeolite-like bridging SiOHAl acid sites in short-range order have also been proposed by theoretical calculations for the strong BASs in ASAs.¹⁰ However, no direct spectral evidence has been reported so far for these strong BASs owing to their low density and their potentially broad signal due to their hydrogen-bond donor nature. Therefore, whether strong BASs are formed via PBS or SiOHAl, or both, is still strongly under debate.

Identifying the surface acid sites and their structure-property relationship is crucial for the rational design and performance optimization of ASAs. Various diffraction techniques, particularly X-ray diffraction (XRD), are powerful for studying the structure of crystalline zeolites. However, they fail to determine unambiguously the local structures of ASAs due to their absence of long-range ordering, resulting in strong line broadenings. IR spectroscopy could identify the surface sites but does not allow distinguishing between PBS and SiOHAl sites. Solid-state NMR spectroscopy is a powerful tool for characterizing surface acid sites and their local structures on silica-alumina materials.^{19, 23} In ¹H 1D spectra under magic-angle spinning (MAS), NMR signals at $\delta_{\text{1H}} = 1.2\text{-}2.2$ ppm for PBS and terminal SiOH groups have been observed.^{17, 24-25} However, the signals at $\delta_{\text{1H}} = 3.6\text{-}5.2$ ppm, typical for bridging SiOHAl acid sites in zeolites,¹⁹ have never been detected so far. Indeed, even when suppressing the strong SiOH signal with a ¹H-^{{27}Al} TRAPDOR (transfer of population in double resonance) filter,²⁶⁻²⁷ only the PBS signal at $\delta_{\text{1H}} \approx 1.8$ ppm has been probed.²⁴

The SENS-DNP technique (surface-enhanced NMR spectroscopy using dynamic nuclear polarization) can significantly enhance the ¹H sensitivity by one or two orders of magnitude.²⁸⁻²⁹ It has been used on ASAs, but still, no SiOHAl sites could be detected.¹⁴ The analysis of the dephasing curves obtained from DNP-enhanced ¹⁷O-^{{27}Al} TRAPDOR experiments have shown that in ASAs the average distances between Al and neighboring SiOH groups are well in line with the theoretical value of PBS.¹⁴ In the same work, the signal at $\delta_{\text{1H}} = 5$ ppm besides that of PBS at $\delta_{\text{1H}} = 1.3$ ppm correlates with ¹⁷O nuclei with slower de-phasing than in ¹⁷O-^{{27}Al} TRAPDOR experiments. This hints at a significantly longer distance between Al-O for the OH signal at $\delta_{\text{1H}} = 5$ ppm, and hence, it was assigned to hydrogen-bonded

OH groups other than bridging OH groups. Therefore, the direct observation of SiOHAl on ASAs remains a great challenge.

²⁷Al-^{{1}H} D-HMQC (Dipolar-mediated hetero-nuclear multiple-quantum correlation between ²⁷Al and ¹H with ²⁷Al observation) two-dimensional (2D) NMR experiments provide information on ¹H-²⁷Al proximities.^{15, 30} This experiment has notably allowed the observation of PBS sites at $B_0 = 18.8$ T on the surface of ASAs. Recently, an ²⁷Al-^{{1}H} D-HMQC experiment at $B_0 = 20.0$ T has been applied to detect the acid sites of aluminum oxide films grown on a silica support.³¹ A very broad signal centered at $\delta_{\text{1H}} = 4.3$ ppm was obtained from the simulation of the indirect ¹H dimension of the ²⁷Al-^{{1}H} D-HMQC spectra, but the indirect ¹H dimension of the corresponding ²⁹Si-^{{1}H} spectra shows only a strong signal at $\delta_{\text{1H}} = 2.5$ ppm without a signal at $\delta_{\text{1H}} = 4.3$ ppm. Therefore, it is still under debate whether the broad signal at $\delta_{\text{1H}} = 4.3$ ppm is due to the presence of bridging SiOHAl groups or mainly a combination of various bridging AlOHAl groups.³²

Here, we report that this experiment carried out at the ultra-high magnetic field of $B_0 = 35.2$ T provides direct evidence for the presence of bridging SiOHAl acid sites on the surface of ASAs. To the best of our knowledge, these sites have never been observed directly so far. This detection is possible since higher B_0 fields increase the resolution along both ¹H and ²⁷Al dimensions and thus allow identifying sites that are undetectable at lower magnetic fields.³⁰ This is particularly the case of ²⁷Al nuclei bonded to bridging OH groups, which are broadened by large quadrupole interactions and hence, which are largely 'invisible' at magnetic fields up to 18.8 T.^{27, 33} Conversely, the reported ²⁷Al-^{{1}H} D-HMQC spectra at 35.2 T allow resolving a new correlation signal between tetra-coordinated Al (Al^{IV}) species ($\delta_{\text{27Al}} = 55\text{-}58$ ppm) and bridging OH groups ($\delta_{\text{1H}} = 4.6$ ppm), demonstrating the presence of bridging silanol groups linked to Al. A similar correlation is also firstly observed with penta-coordinated Al (Al^V) species ($\delta_{\text{27Al}} = 31\text{-}35$ ppm), revealing the formation of SiOHAl BASs based on Al^V species, which have never been probed hitherto. Moreover, we demonstrate that the SiOHAl formation and the distribution of Al species in ASAs depend on the Al/Si ratio. The detailed insights into surface acid sites and the discovery of new SiOHAl in ASAs pave the way for generating strong SiOHAl in ASAs, which facilitates the optimization of their performance for acid catalysis and multifunctional catalysis in various fields.

Amorphous silica-alumina (ASA) catalysts were prepared within microseconds at extremely high temperature (ca. 2000 K) by the flame-spray pyrolysis (FSP) technique.³⁴ Aluminum(III) acetylacetonate (99%, ABCR), tetraethoxysilane (99%, Fluka) acetic acid (analytical grade, Fluka), and methanol (analytical grade, Fluka), were used as received. In brief, the appropriate amounts of the Al and Si precursor materials were dissolved in a 1:1 (vol.) mixture of acetic acid and methanol. The resulting solution was filtered using a

glass filter, pumped through a capillary at a rate of 5 mL.min⁻¹, nebulized by using an O₂ flow rate of 5 L.min⁻¹. The resulting spray was ignited by an annular supporting methane/oxygen flame (1.5/0.9 L.min⁻¹) to generate ASA nanoparticles with a large amount of Al^{IV} species. The ASA samples with different atomic ratios of Al/Si are designated as ASA/X, where X represents the percentage of Al atoms with respect to the total amount of Al and Si atoms in the precursors. The nitrogen isotherms, specific surface areas, concentrations of OH groups, and the densities of BAS and Lewis acid sites (LAS) of these ASA materials were determined in our previous works.^{5, 34-35}

All 2D NMR experiments were performed on the 35.2 Tesla (¹H 1.5 GHz) Series-Connected Hybrid (SCH) magnet operated at the National High Magnetic Field Laboratory (NHMFL, Tallahassee, USA), equipped with a Bruker Avance NEO console. A 2.0 mm triple resonance probe designed and built at the NHMFL was used in double resonance mode and with a sample spinning frequency of $\nu_R = 30$ kHz. The ¹H and ²⁷Al isotropic chemical shifts were referenced indirectly by using a D₂O sample at room temperature (¹⁷O $\delta_{iso} = -2.8$ ppm) and the IUPAC reference sample frequencies in Ref. 36.

The ²⁷Al Triple-Quantum Magic-Angle Spinning (3QMAS) spectra of ASAs were recorded with a shifted-echo pulse sequence,³⁷ and sheared with the QSHEAR program using Matlab.³⁸ Excitation and reconversion pulses lasted $t_p = 3.75$ and 1.25 μ s, respectively, with a radiofrequency (rf) field strength of $\nu_1 = 96$ kHz, and central-transition (CT) selective $\pi/2$ and π pulses of 10 and 20 μ s at $\nu_1 = 8.3$ kHz. The 2D 3QMAS spectra were acquired with complex acquisition using $\Delta t_1 = 33.33$ μ s, $NS = 1920$ transients, 20 t_1 increments, and a recovery delay $T_{RD} = 0.1$ s.

Proximities between ²⁷Al nuclei were probed using 2D ²⁷Al through-space double-quantum to single-quantum (DQ-SQ) homo-nuclear correlation experiments using $\pi/2$ and π -pulses of 10 and 20 μ s CT-selective with $\nu_1 = 8.3$ kHz. The ²⁷Al two-spin DQ coherences were excited and reconverted by applying the BR2₂¹ pulse sequence,³⁹ which reintroduces the ²⁷Al-²⁷Al dipolar interactions under MAS. The lengths of the excitation and reconversion times were optimized on ASA/50 as $\tau_{exc} = \tau_{rec} = 800$ μ s. The rf amplitude applied during the BR2₂¹ pulse sequence was 5 kHz, which corresponds to a nutation frequency of 15 kHz for the ²⁷Al CT, which matches the HORROR (homo-nuclear rotary resonance) condition. Furthermore, a WURST (wideband, uniform rate, smooth truncation) adiabatic pulse with a frequency sweep equal to the MAS frequency and an offset of 300 kHz with respect to the CT was applied before the BR2₂¹ excitation,⁴⁰ to enhance the ²⁷Al CT polarization by inverting/saturating the satellite transitions.⁴¹⁻⁴² The 2D ²⁷Al DQ-SQ spectra resulted from averaging 2048 transients for SA/10 and SA/30, and 3072 transients for ASA/50, respectively, with $\Delta t_1 = 33.33$ μ s, 20 t_1 increments, $T_{RD} = 0.1$ s, resulting in a total experimental time of 2.5 and 3.75 h, respectively.

To enhance the sensitivity, the ²⁷Al nuclei were observed since its rapid longitudinal relaxation allows the use of short recovery delays. The ¹H-²⁷Al dipolar couplings were reintroduced by applying an SR4₁² recoupling on the ¹H channel.⁴³ The ¹H rf amplitudes for the 90° pulses and the SR4₁² recoupling were equal to $\nu_1 \approx 60$ kHz. The CT-selective pulse lengths on ²⁷Al were equal to 10 and 20 μ s for $\pi/2$ and π pulses, respectively, i.e. an rf-field amplitude $\nu_1 = 8.3$ kHz. The total dipolar recoupling time was optimized at $\tau_{rec} = 3.2$ ms. The 2D spectra resulted from the accumulation of $NS = 1024$ transients with States-TPPI acquisition of 16 complex t_1 increments with $\Delta t_1 = 33.33$ μ s and $T_{RD} = 0.1$ s, i.e. a total experiment time of about 1 h.

The formation of BASs strongly depends on the coordination number of Al species in dehydrated ASAs, which could be determined at 35.2 T from 1D MAS (Fig. S1 in Supporting Information) and 2D 3QMAS ²⁷Al NMR spectra (Fig. 1). The 2D 3QMAS spectra of these catalysts are dominated by the Al^{IV} and Al^V signals, while a weaker signal ascribed to hexa-coordinated Al sites (Al^{VI}) is also visible, in agreement with previous reports.^{15, 44} Then the simulation of corresponding 1D ²⁷Al MAS NMR spectra were performed. The obtained quadrupole parameters and concentration of each Al species are summarized in Table S1. With increasing Al/Si ratio from 10/90 to 30/70 and 50/50, called ASA/10, /30 and /50, the content of Al^V is remarkably raised. The high temperature and fast cooling rate under oxygen-rich atmosphere could promote the generation of intermediate Al^V species in ASAs, which may be caused by the local structure reformation under the harsh synthesis conditions.^{3, 24, 45}

Unlike bridging OH groups on zeolites,⁴⁶ only a slight decrease of quadrupole coupling constant (C_Q) was detected for ASA/50 before and after ammonia loading (Table S1). A direct comparison of their 1D ²⁷Al MAS NMR spectra (Fig. S1e) shows an increase of the ²⁷Al signal intensity with small line narrowing. In zeolites, a significant line narrowing for bridging OH groups could be observed even at magnetic field of 9.4 T. This reduced effect is attributed to the wide chemical shift distribution of Al sites in ASA compared to those in crystalline zeolites. Furthermore, for PBS on ASAs, the ammonia loading has limited effect on the ²⁷Al spectra recorded at magnetic field of 16.4 T,¹⁵ which is attributed to the much longer Al-OH distance than in zeolites. The effect upon ammonia loading of the possible bridging SiOHAl groups in ASA could be masked by the signals of other Al sites. Therefore, ²⁷Al-{¹H} D-HMQC 2D NMR experiments were applied to probe selectively surface protons in connection with Al sites.

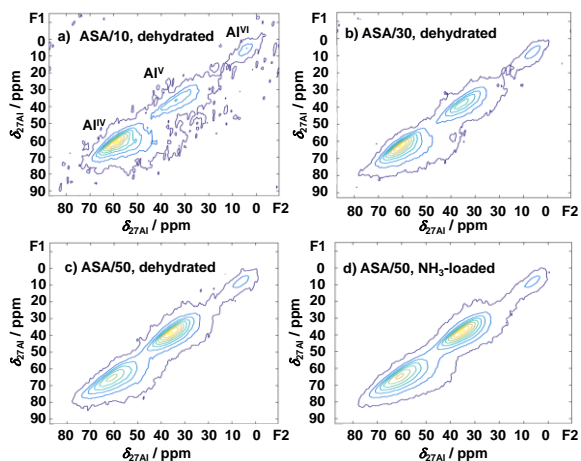


Figure 1. ^{27}Al 3QMAS ASA spectra recorded at 35.2 T with $\nu_R = 30$ kHz of (a) ASA/10, (b) ASA/30, and (c) ASA/50, dehydrated at 723 K under vacuum for 12 h, as well as (d) ASA/50 loaded with ammonia and evacuated at 373 K for 1 h.

In SA10, the dominant Al^{IV} and Al^{V} species are highly dispersed in the silica networks on the surface,⁴⁴ which is proposed to promote the formation of moderate BASs.¹⁵ $^{27}\text{Al}\{-^1\text{H}\}$ D-HMQC 2D spectra at $B_0 \leq 18.8$ T only exhibit a broad ^{27}Al signal at 60–20 ppm correlated to the ^1H signal at 1.9 ppm (see Fig. S2). When the width of the ^{27}Al resonance is dominated by the second-order quadrupolar interaction, the resolution of the ^{27}Al 1D spectra is proportional to $(B_0)^2$, and hence is improved at higher magnetic fields. Here, cross-peaks between both Al^{IV} ($\delta_{27\text{Al}} = 55$ ppm) and Al^{V} ($\delta_{27\text{Al}} = 31$ ppm) species interacting with neighboring SiOH groups ($\delta_{\text{H}} = 1.9$ ppm) are clearly resolved at $B_0 = 35.2$ T (Fig. 2a and 2b), and confirmed by the corresponding ^{27}Al 1D slices (Fig. 2d and 2e). These cross-peaks are ascribed to PBSs generated in ASA/10 (Fig. 2f), responsible for the moderate Brønsted acidity of ASAs with low Al/Si ratios. Moreover, an additional weak ^1H signal at $\delta_{\text{H}} = 3.4$ ppm correlates with both Al^{IV} and Al^{V} species (Fig. 2a and 2c) has been assigned to AIOH groups.¹⁹ Since terminal AIOH groups are often observed at δ_{H} shifts ranging from -0.5 to 0.7 ppm, the strong observed deshielding indicates that these AIOH are either polarized by neighboring sites or bridged to another Al sites, as in doubly bridging (μ_2) AIOHAl sites.^{19, 47} The ^{27}Al DQ-SQ 2D spectrum of ASA/10 (Fig. S3a) exhibits $\text{Al}^{\text{IV}}\text{-Al}^{\text{IV}}$, $\text{Al}^{\text{IV}}\text{-Al}^{\text{V}}$, and $\text{Al}^{\text{V}}\text{-Al}^{\text{V}}$ correlations, which were undetectable in previous works at 18.8 T due to their low concentration and strong line broadening.³ Since no strong BASs could be detected in ASA/10,³⁴ these proximities between Al sites are not accompanied by a significant synergistic effect on the same SiOH group.³ The detected proximities between Al sites are proposed to be existing in a few nearby Al species, requiring AIOH groups to balance the local structure as supported by DFT calculation studies.⁴⁴ These AIOH groups produce the deshielded ^1H signal at $\delta_{\text{H}} = 3.4$ ppm. In addition to the $\text{SiOH}\cdots\text{Al}^{\text{IV}}$ sites, the presence of $\text{Al}^{\text{IV}}\text{-O-SiOH}$ species

has been constructed by DFT calculations from the silication of the (100) $\gamma\text{-Al}_2\text{O}_3$ surface,⁴⁸ which could be centred at $\delta_{\text{H}} = 1.9$ ppm as well. After loading with NH_3 , the ^1H peak of acidic $\text{SiOH}\cdots\text{Al}^{\text{IV}}$ sites is shifted to ca. 6.7 ppm for NH_4^+ , while the ^1H peak of non-acidic $\text{Al}^{\text{IV}}\text{-O-SiOH}$ species remains at around 1.9 ppm. Upon NH_3 loading, the ^1H peak at 1.9 ppm in Fig. 2a was disappeared and replaced by a peak at $\delta_{\text{H}} = 6.7$ ppm,¹⁵ demonstrating the absence of $\text{Al}^{\text{IV}}\text{-O-SiOH}$ species.

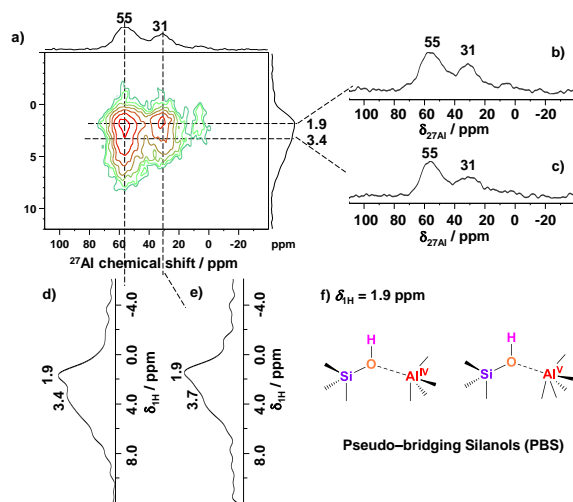


Figure 2. Correlation spectra of dehydrated ASA/10. (a) 2D $^{27}\text{Al}\{-^1\text{H}\}$ D-HMQC spectrum acquired using $\tau_{\text{rec}} = 3.2$ ms, along with the ^{27}Al and ^1H 1D slices displayed in subfigures (b,c) and (d,e), respectively. (f) Proposed structure of PBSs corresponding to the dominant cross-peaks in the $^{27}\text{Al}\{-^1\text{H}\}$ D-HMQC spectrum. ASA/10 was dehydrated at 723 K for 12 h under vacuum and the spectra were recorded at 35.2 T with $\nu_R = 30$ kHz.

At higher Al concentration, besides the signals of moderate BASs and AIOH groups at $\delta_{\text{H}} = 1.9$ and 3.4 ppm, an additional signal at $\delta_{\text{H}} = 4.6$ ppm was detected in dehydrated ASA/30 (Fig. 3a). Physisorbed and hydrogen-bonded water molecules on ASAs show ^1H signals at ca. $\delta_{\text{H}} = 4.6\text{--}4.8$ ppm, which could be probed by ^1H MAS NMR even at low magnetic fields (7.0 T).⁴⁹ However, these signals are undetectable in ^1H DQ-SQ 2D spectra at 18.8 T for dehydrated ASAs as used here.³ Hence, the presence of water molecules with signals at $\delta_{\text{H}} = 4.6$ ppm can be ruled out. The formation of bridging AIOH groups at similar ^1H chemical shifts ($\delta_{\text{H}} = 3\text{--}4.5$ ppm) has been reported recently for Al_2O_3 at 35.2 T.⁴⁷ On the other hand, the signal at $\delta_{\text{H}} = 4.6$ ppm is typical for bridging OH groups in crystalline zeolites.³²

In ASAs, BASs originate from Al^{IV} or Al^{V} sites interacting with neighboring SiOH groups in the alumina-silica interphase region or inside the silica networks, whereas Al^{VI} species are typically in the alumina phase and do not contribute to the formation of BASs. The ^{27}Al DQ-SQ 2D spectrum of dehydrated ASA/30 (Fig. S3b) does not exhibit any correlation with Al^{VI} signal, suggesting that no significant alumina phase

exists in ASA/30, which is in line with a previous report.⁴⁴ The *D*-HMQC spectrum shows cross-peaks between Al^{VI} species ($\delta_{27\text{Al}} = 7$ ppm) and either SiOH ($\delta_{1\text{H}} = 1.9$ ppm) or bridging AlOH ($\delta_{1\text{H}} = 3.4$ ppm) groups (see the slices shown in Figs. 3b, 3c and 3g). The NMR observations suggest that some Al^{VI} species are isolated on the silica surface, whereas other ones are located in the proximity of AlOH groups in Figs. 3c and 3g. However, the Al^{VI} species are not in connection with the signal at $\delta_{1\text{H}} = 4.6$ ppm in ASA/30 (Fig. 3d), demonstrating that this signal is not associated with the alumina phase. The signal at $\delta_{1\text{H}} = 4.6$ ppm mainly correlates with Al^{IV} or Al^V signals, and thus, is considered to be located either in the alumina-silica interphase region or inside the silica networks, contributing to BASs (Fig. 3h).

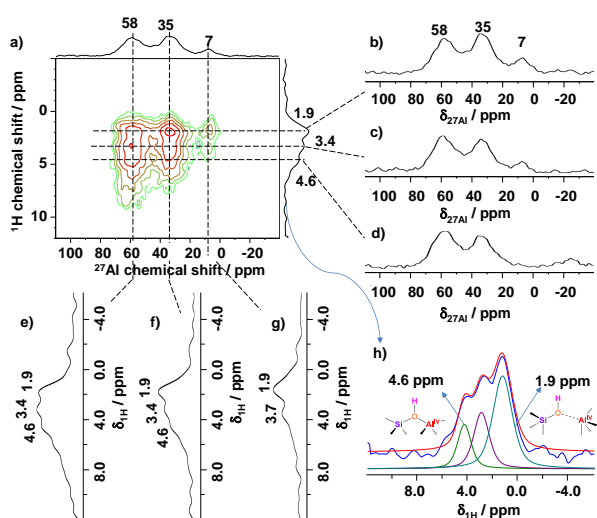


Figure 3. Correlation spectra of dehydrated ASA/30. (a) $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC 2D spectrum acquired using $\tau_{\text{rec}} = 3.2$ ms, along with the ^{27}Al and ^1H 1D slices displayed in subfigures (b-d) and (e-g), respectively. (h) Simulation of the ^1H projection of (a) and the proposed local environments of silanol groups based on spectrum (a) and corresponding slices. ASA/30 was dehydrated at 723 K for 12 h under vacuum and spectra were recorded at 35.2 T with $\nu_{\text{R}} = 30$ kHz.

Similar structures are also detected for ASAs with higher Al/Si ratios (Fig. 4). However, the signals at $\delta_{1\text{H}} = 4.6$ and 3.4 ppm are overlaid by the intense one at $\delta_{1\text{H}} = 1.9$ ppm in the $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC 2D spectrum of dehydrated ASA/50 (Fig. 4a). The cross-peaks between Al^{IV} and Al^V sites and the ^1H signal at 1.9 ppm are ascribed to PBSs interacting with one or two nearby Al atoms.^{3, 15} The first case leads to moderate BASs, which can catalyze the dehydration of cyclohexanol.⁶ However, we have recently shown that the presence of an additional Al atom in the vicinity of PBSs can enhance their Brønsted acidity.³ This increased intensity of the ^1H signal at 1.9 ppm at high Al/Si ratio indicates an increased amount of PBSs, which is consistent with a homogeneous distribution of Si atoms into the alumina

matrix or on the alumina surface of ASAs at high Al/Si ratios.¹³⁻¹⁴

The *D*-HMQC 2D spectrum of dehydrated ASA/50 shown in Fig. 4a exhibits cross-peaks between Al^{IV} and Al^V sites and the ^1H signal at $\delta_{1\text{H}} = 5.2$ ppm. This ^1H chemical shift has been ascribed to aluminols, bridging OH groups, or hydrogen-bonded hydroxyls.¹⁹ Similar to bridging AlOH groups observed at $\delta_{1\text{H}} = 3.4$ ppm (Figs. 2c, 3c, and S4b), the OH groups at $\delta_{1\text{H}} = 5.2$ ppm is in proximity to Al^{VI} species, but the chemical shift is significantly higher than that of triply-bridging AlOH groups in alumina resonating at $\delta_{1\text{H}} = 3.4$ ppm.⁴⁷ Hence, the cross-peaks at $\delta_{1\text{H}} = 5.2$ ppm could be ascribed to triply-bridging AlOH in strong electrostatic interaction with neighboring oxygen atoms, similar to hydrogen-bonded hydroxyls observed in zeolites.¹⁹

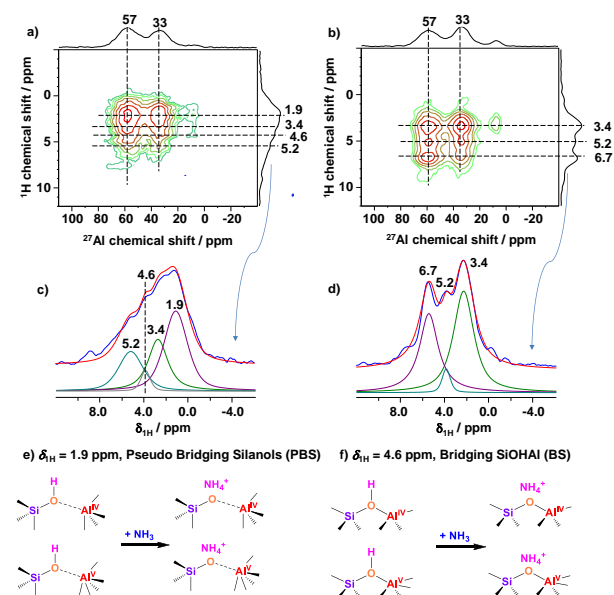


Figure 4. $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC spectra of ASA/50: (a) dehydrated at 723 K for 12 h under vacuum, and (b) upon ammonia loading. (c,d) Simulation of the ^1H projection of (a) and (b). (e,f) Proposed structures of PBS and BS sites before and after their reaction with ammonia. Spectra were recorded at 35.2 T with $\nu_{\text{R}} = 30$ kHz and $\tau_{\text{rec}} = 3.2$ ms.

The acidic properties of surface protons could be identified using NH_3 as probe molecule because the protonation of NH_3 at BASs results in NH_4^+ ions producing a ^1H signal at $\delta_{1\text{H}} = 6.5\text{--}7.0$ ppm, which can clearly be distinguished from other OH groups in silica-aluminas. NH_3 coordinated to Lewis acidic Al sites on Al_2O_3 and ASA leads to ^1H signals in the range $\delta_{1\text{H}} = -0.5 \leftrightarrow 4.6$ ppm.^{35, 50-52} AlOH groups are generally inactive in protonating NH_3 . Although a recent report shows protonation of NH_3 on Al_2O_3 , the amount of potential acidic OH groups is very small compared to LASs.⁵⁰ Since AlOH groups are inactive for the protonation of ammonia, the increase in the signal intensity of the cross-peaks between Al^{IV} resonance and ^1H signals at 3.4 and 5.2 ppm upon ammonia loading is attributed to the adsorption of ammonia at LASs (Figs. S4b,g and S4d,h).³⁵ In the $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC 2D spectrum shown in Fig. 4b, the cross-peaks at $\delta_{1\text{H}} = 1.9$ and 4.6 ppm, have

disappeared compared to Fig. 4a, while concomitantly a signal at $\delta_{\text{IH}} = 6.7$ ppm appears. The large intensity of the ammonium signal and the very limited change in the signals of AlOH groups (Figs. S4b,g and S4d,h) indicate that the proton donation arises mainly from the peaks at $\delta_{\text{IH}} = 1.9$ and 4.6 ppm. These spectral modifications demonstrate that these two proton species are surface BASs. The complete deprotonation (disappearance) of the peak at $\delta_{\text{IH}} = 1.9$ ppm in Fig. 4b is because that mainly acidic SiOH groups with neighboring Al centers can be selectively detected in $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC experiments, while non-acidic SiOH groups are still observed upon ammonia adsorption using conventional ^1H 1D experiments (Fig. S5).^{15, 34} The signal at $\delta_{\text{IH}} = 1.9$ ppm is widely accepted for PBS sites in ASAs. DFT calculations shows that the interaction between SiOH groups and local Al sites in ASAs is very weak compared to bridging SiOHAl groups, unless deprotonation upon the adsorption of probe molecules.⁵³ Therefore, both moderate and strong PBS sites denote similar signals at $\delta_{\text{IH}} = 1.9$ ppm, without strong deshielding on the nearby silanol protons. Therefore, the signal at $\delta_{\text{IH}} = 4.6$ ppm does not arise from the PBS and the bridging AlOH in alumina. The acidic hydroxyl groups with $\delta_{\text{IH}} = 4.6$ ppm can correlate with Al^{IV} and Al^V sites (Figs. S4c,e,f,j,k). In zeolites, Al^{IV} sites lead to a significant deshielding of the ^1H resonance on neighboring OH groups, like in bridging SiOHAl^{IV} groups. Therefore, the cross peaks between the signal at $\delta_{\text{IH}} = 4.6$ ppm and Al^{IV} and Al^V peaks are ascribed here to bridging SiOHAl^{IV} and SiOHAl^V moieties. These two groups transfer protons to ammonia. To the best of our knowledge, SiOHAl^V acid sites have never been observed so far.

As the hetero-nuclear coherence transfer in $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC experiments are only effective up to a few Å, the ^1H signal of ammonium ions only correlates with ^{27}Al signals involved in the conjugated base of BASs. Hence, the cross-peaks of the ^1H signal of ammonium ions in *D*-HMQC spectra demonstrate that BASs contain both Al^{IV} and Al^V sites. The structures of PBS and BS involving Al^{IV} and Al^V sites and their reactions with ammonia are shown in Figs. 4e and 4f. Both PBS and BS act as BASs. To determine the population densities of BASs based on PBS and BS, we assume that they have comparable transfer efficiencies in the *D*-HMQC experiments. Then the population densities of PBS and BS are determined from the integrated intensities of their signal in the ^1H projection of *D*-HMQC spectra (Figs. 3h, 4c and 4d), and the as-derived molar fractions are given in Table 1. The fraction of SiOHAl amounts to the range of 9.2-24 % of the total number of BASs, and PBS are predominant in ASAs. With increasing Al/Si ratio from 30/70 to 50/50, the molar fraction and the number of SiOHAl were significantly reduced.

Table 1. Molar fraction and concentration of BS and PBS BASs on ASAs.^a

	Molar fraction of BS ^b / %	Density of BS / mmol/g	Density of PBSs / mmol/g
ASA/10	0	0	9.8×10^{-2}
ASA/30	18.3	2.0×10^{-2}	9.1×10^{-2}
ASA/50	9.2	1.2×10^{-2}	12.2×10^{-2}

^aTotal densities of BASs in ASAs are obtained from ref.34. ^bMolar fractions of BS were determined by the ratios (BS)/(BS+PBS) of the integrated intensities obtained from simulations of the ^1H projections of $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC 2D spectra of dehydrated ASA.

Bridging OH groups are widely accepted to be generated based on framework Al^{IV} species in crystalline zeolites, but are rarely probed in ASAs. Generally, Al^{IV} species involved in SiOHAl^{IV} groups are subject to very large C_Q values, e.g. up to 15 ± 1 MHz for framework Al^{IV} species in zeolite H-Y,³³ causing significant line broadenings, and thus are ‘invisible’ at low to medium magnetic fields. Increasing the magnetic field strength and the spinning rate may significantly narrow the ^{27}Al signal, and allow recovering this ‘invisible’ signal. Moreover, $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC 2D spectra allow the selective detection of SiOH groups with different neighboring Al sites, while removing the ^1H signal of non-acidic terminal SiOH groups.

With this in mind, the newly observed signals at $\delta_{\text{IH}} = 4.6$ ppm in proximity to Al^{IV} species demonstrate the formation of SiOHAl groups in ASAs, which are similar to those observed in zeolites with the same $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC 2D experiments.³⁰ These ^{27}Al and ^1H correlations are undetectable when using a magnetic field of 18.8 T (Fig. S2), but could be detected at 35.2 T. This observation confirms that these Al species are subject to large quadrupolar interactions. Large values of $C_Q = 13$ -16.6 MHz, similar to those reported for framework Al^{IV} species, have been calculated in ASAs by DFT for Al^{IV} with neighboring silanols.⁴⁴ These SiOHAl environments of ASAs exhibit similar ^1H chemical shifts and acidic properties as bridging OH groups in zeolites.

More interestingly, Al^V species exhibit similar proximity to bridging OH groups at $\delta_{\text{IH}} = 4.6$ ppm, as observed with Al^{IV} species. Al^V species interacting with SiOH groups can exhibit $C_Q \approx 12$ -16 MHz, according to DFT calculations,⁴⁴ similar to those reported for framework Al^{IV} species in zeolite H-Y.³³ $^{27}\text{Al}\{-^1\text{H}\}$ *D*-HMQC 2D spectra, recorded before and after ammonia loading (Figs. 4a and 4b), demonstrate that Al^V species are in proximity to protons at $\delta_{\text{IH}} = 4.6$ ppm and ammonium ions (Figs. S4f,k), and are thus acting as SiOHAl. The cross-peaks with Al^V signal in *D*-HMQC 2D spectra have intensities similar to those with Al^{IV} resonance (Figs. 3d and S4c). Therefore, bridging sites interacting with Al^V sites represent a significant fraction of the total amount of the bridging sites and not only some extra sites in the structure. Furthermore, the reaction of ASAs with ammonia (Fig. 4b) indicates that Al^{IV} and Al^V are unsaturated species with similar

properties to promote the acid strength of neighboring silanols ($\delta_{\text{H}} = 1.9$ ppm).

The formation of bridging SiOHAl groups in ASAs has been proposed by DFT calculation studies, constructed from the silication of the (100) γ -Al₂O₃ surface upon hydrothermal treatment.^{10, 13, 53-54} The presence of bridging SiOHAl groups requires a high hydroxyl coverage (≥ 5.4 OH nm⁻²),¹³ much higher than that (ca. 2.8~3.0 OH nm⁻²) on flame-made ASAs.³⁴ The surface density of SiOHAl groups in ASA/30 is ca. 0.5 nm⁻² at 2.8 OH nm⁻², which is similar to that reported for DFT model at 7.5 OH nm⁻², indicating a promotion of SiOHAl formation using FSP. In the DFT model, the bridging SiOHAl groups are formed based on Al^V species, other than the well accepted Al^{IV} species in zeolites.⁵⁴ However, SiOHAl groups based on both Al^{IV} and Al^V species are observed in this work. The DFT obtained structural parameters of SiOHAl groups in ASAs are like those reported for zeolites, such as bond distances (Al-O, Si-O and O-H) and Si-O-Al angle.¹³ With the similar structure, similar large C_Q values of SiOHAl in ASAs and zeolites can be expected and are reported in DFT calculation studies as discussed above.⁴⁴ Therefore, our observations of SiOHAl groups in ASAs can be supported by the existing DFT calculation studies.

To the best of our knowledge, the bridging silanol groups SiOHAl^V have never been reported experimentally so far and the above results represent the first direct observation of bridging silanol groups in ASAs, in which a PBS local environment dominates. These observations well support the strong BASs probed in ASAs in previous works.²⁰⁻²² The formation of BS could depend on the wide distribution of Al species in ASAs, leading to BASs with versatile strengths and structures. For ASAs prepared at milder conditions, mainly moderate BASs based on Al^{IV} sites are generated, and thereby observing directly the structure of a few strong BASs is quite difficult. Likewise, distinguishing the structure of strong BASs from PBS with an Al-synergistic effect and a BS site is impossible using other techniques, such as FTIR. In flame-spray pyrolysis, the high flame temperature (~ 2000 K) and oxygen-rich atmosphere promote the formation of Si-O-Al bonds, leading to highly populated surface Al^V species and much higher BAS density than in ASAs prepared at milder conditions.^{24, 44} The harsh synthesis conditions using FSP provide a more homogeneous distribution of Al and Si atoms even at high Al loading, which tends to promote the formation of bridging silanol groups based on either Al^{IV} or Al^V sites in flame-derived ASAs. Hence, we could directly observe here that the formation of bridging SiOHAl is not a unique property of zeolites featuring long-range order, but is also possible in amorphous materials, such as ASAs.

As the formation and the amount of SiOHAl sites depend on the Al/Si ratio and Al distribution, it may be possible to tune the acidity of ASAs for the selective production of middle distillates from hydrocracking. In addition, the strong SiOHAl BASs can cooperate with

other sites on the surface of ASAs for the catalysis of multi-step reactions, including bifunctional Brønsted-Lewis acid-catalytic glucose dehydration for 5-hydroxymethylfurfural.⁵⁵⁻⁵⁷ Moreover, bridging OH groups could result in strong metal-support interactions,⁵⁸ enhancing the catalytic activity and stability of active metal species for removing hazardous gases, such as NO, NO₂, NH₃, and volatile organic compounds,⁵⁹⁻⁶⁰ with benefits to the environment and human health. Therefore, our discovery of SiOHAl formation in ASAs could significantly expand their applications in various fields.

In summary, the formation of Al^{IV}-based bridging OH groups in ASAs, together with a new type of Al^V-based bridging OH groups, are directly observed by ²⁷Al-{¹H} *D*-HMQC 2D NMR spectroscopy. So far, bridging OH groups, as the key catalytic active sites, have been widely accepted to be formed with framework Al^{IV} species in crystalline zeolites. Compared to zeolites, the formation of bridging OH groups in ASAs is strongly under debate due to the lack of long-range order in these materials, longer Al-O distance, significantly lower Brønsted acidity, and rare spectral evidence. Here, we proved the formation of bridging OH groups in ASAs based on Al^{IV} and Al^V species in proximity to SiOH groups, showing similar ¹H isotropic chemical shifts and acidic properties as those in zeolites. We demonstrated that the formation of bridging OH groups is general in silica-aluminas and virtually independent from the long-range order of the local structure. This could be achieved by reducing the distance between Al^{IV}/Al^V species and SiOH groups in compact structure in ASAs by using a suitable synthesis approach. Our findings, which are related to the use of a very high magnetic field (35.2 T), show promise for the development of ASAs populated with bridging OH groups, enhancing their acidity and catalytic activity for a wide range of acid and multifunctional catalysis for energy and chemical production in key industrial chemical processes.

ASSOCIATED CONTENT

The Supporting Information is available free of charge via the Internet at <http://pubs.acs.org>.

Method for simulating the 1D ²⁷Al MAS NMR spectra

1D ²⁷Al and ¹H MAS NMR spectra

²⁷Al-{¹H} *D*-HMQC 2D spectrum of ASA/10

²⁷Al and ¹H DQ-SQ 2D spectrum of ASAs

²⁷Al and ¹H slices from Fig. 4

XRD patterns, STEM and EDX images of ASAs

Nitrogen adsorption/desorption isotherms of ASAs

Table S1. Summary of parameters and deconvolution results

Table S2. Surface areas and acidic properties of ASA catalysts

AUTHOR INFORMATION

Corresponding Author

* olivier.lafon@univ-lille.fr; jun.huang@sydney.edu.au

Author Contributions

All authors have given approval to the final version of the manuscript.

Notes

The authors declare no competing financial interest.

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REFERENCES

- (1) Li, Y.; Li, L.; Yu, J. H. Applications of Zeolites in Sustainable Chemistry. *Chem.* **2017**, *3*, 928-949.
- (2) Corma, A. Inorganic solid acids and their use in acid-catalyzed hydrocarbon reactions. *Chem. Rev.* **1995**, *95*, 559-614.
- (3) Wang, Z.; Li, T.; Jiang, Y.; Lafon, O.; Liu, Z.; Trébosc, J.; Baiker, A.; Amoureux, J.-P.; Huang, J. Acidity enhancement through synergy of penta- and tetra-coordinated aluminum species in amorphous silica networks. *Nat. Commun.* **2020**, *11*, 225.
- (4) Wang, Z.; Pokhrel, S.; Chen, M.; Hunger, M.; Mädler, L.; Huang, J. Palladium-doped silica-alumina catalysts obtained from double-flame FSP for chemoselective hydrogenation of the model aromatic ketone acetophenone. *J. Catal.* **2013**, *302*, 10-19.
- (5) Wang, Z.; Jiang, Y.; Baiker, A.; Huang, J. Efficient acid-catalyzed conversion of phenylglyoxal to mandelates on flame-derived silica/alumina. *ACS Catal.* **2013**, *3*, 1573-1577.
- (6) Mouat, A. R.; George, C.; Kobayashi, T.; Pruski, M.; van Duyne, R. P.; Marks, T. J.; Stair, P. C. Highly dispersed SiO_x/Al₂O₃ catalysts illuminate the reactivity of isolated silanol sites. *Angew. Chem. Int. Ed.* **2015**, *54*, 13346-13351.
- (7) Wang, Y. J.; Liu, C.; Wu, Y. J.; Song, Y. H.; Zhu, M. L.; Huang, J.; Liu, Z. T.; Liu, Z. W. Flame-spray-pyrolysis amorphous alumina-silica for tailoring the product distribution of Fischer-Tropsch synthesis. *Catal. Today*, **2020**, *339*, 40-47.
- (8) Yun, D.; Yun, Y. S.; Kim, T. Y.; Park, H.; Lee, J. M.; Han, J. W.; Yi, J. Mechanistic study of glycerol dehydration on Brønsted acidic amorphous aluminosilicate. *J. Catal.* **2016**, *341*, 33-43.
- (9) Aitani, A. M. Oil Refining and Products. In *Encyclopedia of Energy*; Cleveland, C. J., Eds.; Elsevier: New York, U.S.A. 2004; pp 715-729.
- (10) Chizallet, C. Toward the atomic scale simulation of intricate acidic aluminosilicate catalysts. *ACS Catal.* **2020**, *10*, 5579-5601.
- (11) Busca, G. Silica-alumina catalytic materials: A critical review. *Catal. Today* **2020**, *357*, 621-629.
- (12) Wang, Z.; Buechel, R.; Jiang, Y.; Wang, L.; Xu, H.; Castignolles, P.; Gaborieau, M.; Lafon, O.; Amoureux, J.-P.; Hunger, M.; Baiker, A.; Huang, J. Engineering the distinct structure interface of subnano-alumina domains on silica for acidic amorphous silica-alumina toward biorefining. *JACS Au* **2021**, *1*, 262-271.
- (13) Chizallet, C.; Raybaud, P. Pseudo-bridging silanols as versatile Brønsted acid sites of amorphous aluminosilicate surfaces. *Angew. Chem. Int. Ed.* **2009**, *48*, 2891-2893.
- (14) Perras, F. A.; Wang, Z.; Kobayashi, T.; Baiker, A.; Huang, J.; Pruski, M. Shedding light on the atomic-scale structure of amorphous silica-alumina and its Brønsted acid sites. *Phys. Chem. Chem. Phys.* **2019**, *21*, 19529-19537.
- (15) Wang, Z. C.; Jiang, Y. J.; Lafon, O.; Trebosc, J.; Kim, K. D.; Stampfl, C.; Baiker, A.; Amoureux, J.-P.; Huang, J. Brønsted acid sites based on penta-coordinated aluminum species. *Nat. Commun.* **2016**, *7*, 13820.
- (16) Luo, Q.; Deng, F.; Yuan, Z. Y.; Yang, J.; Zhang, M. J.; Yue, Y.; Ye, C. H. Using trimethylphosphine as a probe molecule to study the acid states in Al-MCM-41 materials by solid-state NMR spectroscopy. *J. Phys. Chem. B* **2003**, *107*, 2435-2442.
- (17) Wang, Z.; Ling, H.; Shi, J.; Stampfl, C.; Yu, A.; Hunger, M.; Huang, J. Acidity enhanced [Al]MCM-41 via ultrasonic irradiation for the Beckmann rearrangement of cyclohexanone oxime to ε-caprolactam. *J. Catal.* **2018**, *358*, 71-79.
- (18) Xu, J.; Wang, Q.; Deng, F. Metal active sites and their catalytic functions in zeolites: Insights from solid-state NMR spectroscopy. *Acc. Chem. Res.* **2019**, *52*, 2179-2189.
- (19) Jiang, Y.; Huang, J.; Dai, W.; Hunger, M. Solid-state nuclear magnetic resonance investigations of the nature, property, and activity of acid sites on solid catalysts. *Solid State Nucl. Magn. Reson.* **2011**, *39*, 116-141.
- (20) Bevilacqua, M.; Montanari, T.; Finocchio, E.; Busca, G. Are the active sites of protonic zeolites generated by the cavities? *Catal. Today* **2006**, *116*, 132-142.
- (21) Xu, B.; Sievers, C.; Lercher, J. A.; van Veen, J. A. R.; Giltay, P.; Prins, R.; van Bokhoven, J. A. Strong Brønsted

- acidity in amorphous silica–aluminas. *J. Phys. Chem. C* **2007**, *111*, 12075-12079.
- (22) Poduval, D. G.; van Veen, J. A. R.; Rigutto, M. S.; Hensen, E. J. M. Brønsted acid sites of zeolitic strength in amorphous silica-alumina. *Chem. Commun.* **2010**, *46*, 3466-3468.
- (23) Xu, J.; Wang, Q.; Li, S.; Deng, F. Solid-state NMR characterization of acid properties of zeolites and solid acid catalysts. In *Solid-State NMR in Zeolite Catalysis*; Springer: Singapore, 2019; pp 159-197.
- (24) Wang, Z.; Jiang, Y.; Jin, F.; Stampfl, C.; Hunger, M.; Baiker, A.; Huang, J. Strongly enhanced acidity and activity of amorphous silica–alumina by formation of pentacoordinated Al^V species. *J. Catal.* **2019**, *372*, 1-7.
- (25) Wang, Z.; Jiang, Y.; Rachwalik, R.; Liu, Z.; Shi, J.; Hunger, M.; Huang, J. One-step room-temperature synthesis of [Al]MCM-41 materials for the catalytic conversion of phenylglyoxal to ethylmandelate. *ChemCatChem.* **2013**, *5*, 3889-3896.
- (26) Deng, F.; Yue, Y.; Ye, C. H., ¹H/²⁷Al TRAPDOR NMR studies on aluminum species in dealuminated zeolites. *Solid State Nucl. Magn. Reson.* **1998**, *10*, 151-160.
- (27) Grey, C. P.; Vega, A. J. Determination of the quadrupole coupling constant of the invisible aluminum spins in zeolite HY with ¹H/²⁷Al TRAPDOR NMR. *J. Am. Chem. Soc.* **1995**, *117*, 8232-8242.
- (28) Perras, F. A.; Kobayashi, T.; Pruski, M. Natural abundance ¹⁷O DNP two-dimensional and surface-enhanced NMR spectroscopy. *J. Am. Chem. Soc.* **2015**, *137*, 8336-8339.
- (29) Liao, W. C.; Ghaffari, B.; Gordon, C. P.; Xu, J.; Copéret, C. Dynamic nuclear polarization surface enhanced NMR spectroscopy (DNP SENS): Principles, protocols, and practice. *Current Opinion in Colloid & Interface Sci.* **2018**, *33*, 63-71.
- (30) Chen, K.; Horstmeier, S.; Nguyen, V. T.; Wang, B.; Crossley, S. P.; Pham, T.; Gan, Z.; Hung, I.; White, J. L. Structure and catalytic characterization of a second framework Al(IV) site in zeolite catalysts revealed by NMR at 35.2 T. *J. Am. Chem. Soc.* **2020**, *142*, 7514-7523.
- (31) Kaushik, M.; Leroy, C.; Chen, Z.; Gajan, D.; Willinger, E.; Müller, C. R.; Fayon, F.; Massiot, D.; Fedorov, A.; Copéret, C.; Lesage, A.; Florian, P. Atomic-scale structure and its impact on chemical properties of aluminum oxide layers prepared by atomic layer deposition on silica. *Chem. Mater.* **2021**, *33*, 3335-3348.
- (32) Yang, W.; Wang, Z.; Huang, J.; Jiang, Y. Qualitative and quantitative analysis of acid properties for solid acids by solid-state nuclear magnetic resonance spectroscopy. *J. Phys. Chem. C* **2021**, *125*, 10179-10197.
- (33) Jiao, J.; Kanellopoulos, J.; Wang, W.; Ray, S. S.; Foerster, H.; Freude, D.; Hunger, M. Characterization of framework and extra-framework aluminum species in non-hydrated zeolites Y by ²⁷Al spin-echo, high-speed MAS, and MQMAS NMR spectroscopy at B-0=9.4 to 17.6 T. *Phys. Chem. Chem. Phys.* **2005**, *7*, 3221-3226.
- (34) Huang, J.; van Vegten, N.; Jiang, Y.; Hunger, M.; Baiker, A. Increasing the Brønsted acidity of flame-derived silica/alumina up to zeolitic strength. *Angew. Chem. Int. Ed.* **2010**, *49*, 7776-7781.
- (35) Wang, Z.; Jiang, Y.; Stampfl, C.; Baiker, A.; Hunger, M.; Huang, J. NMR spectroscopic characterization of flame-made amorphous silica-alumina for cyclohexanol and glyceraldehyde conversion. *ChemCatChem.* **2020**, *12*, 287-293.
- (36) Harris, R. K.; Becker, E. D.; Cabral de Menezes, S. M.; Goodfellow, R.; Granger, P. NMR nomenclature: Nuclear spin properties and conventions for chemical shifts: IUPAC recommendations 2001. *Solid State Nucl. Magn. Reson.* **2002**, *22*, 458-483.
- (37) Massiot, D.; Touzo, B.; Trumeau, D.; Coutures, J. P.; Virlet, J.; Florian, P.; Grandinetti, P. J. Two-dimensional magic-angle spinning isotropic reconstruction sequences for quadrupolar nuclei. *Solid State Nucl. Magn. Reson.* **1996**, *6*, 73-83.
- (38) Hung, I.; Trébosc, J.; Hoatson, G. L.; Vold, R. L.; Amoureux, J.-P.; Gan, Z. Q-shear transformation for MQMAS and STMAS NMR spectra. *J. Magn. Reson.* **2009**, *201*, 81-86.
- (39) Wang, Q.; Hu, B.; Lafon, O.; Trébosc, J.; Deng, F.; Amoureux, J.-P. Double-quantum homonuclear NMR correlation spectroscopy of quadrupolar nuclei subjected to magic-angle spinning and high magnetic field. *J. Magn. Reson.* **2009**, *200*, 251-260.
- (40) Siegel, R.; Nakashima, T. T.; Wasylshen, R. E. Signal enhancement of NMR spectra of half-integer quadrupolar nuclei in solids using hyperbolic secant pulses. *Chem. Phys. Lett.* **2004**, *388*, 441-445.
- (41) Yao, Z.; Kwak, H.-T.; Sakellariou, D.; Emsley, L.; Grandinetti, P. J. Sensitivity enhancement of the central transition NMR signal of quadrupolar nuclei under magic-angle spinning. *Chem. Phys. Lett.* **2000**, *327*, 85-90.
- (42) Kwak, H. T.; Prasad, S.; Clark, T.; Grandinetti, P. J. Enhancing sensitivity of quadrupolar nuclei in solid-state NMR with multiple rotor assisted population transfers. *Solid State Nucl. Magn. Reson.* **2003**, *24*, 71-77.
- (43) Brinkmann, A.; Kentgens, A. P. M. Proton-selective ¹⁷O–¹H distance measurements in fast magic-angle-spinning solid-state NMR spectroscopy for the determination of hydrogen bond lengths. *J. Am. Chem. Soc.* **2006**, *128*, 14758-14759.
- (44) Wang, Z.; Jiang, Y.; Yi, X.; Zhou, C.; Rawal, A.; Hook, J.; Liu, Z.; Deng, F.; Zheng, A.; Hunger, M.; Baiker, A.; Huang, J. High population and dispersion of pentacoordinated Al^V species on the surface of flame-made amorphous silica-alumina. *Sci. Bull.* **2019**, *64*, 516-523.
- (45) Kim, K. D.; Pokhrel, S.; Wang, Z. C.; Ling, H. J.; Zhou, C. F.; Liu, Z. W.; Hunger, M.; Mädler, L.; Huang, J. Tailoring high-performance Pd catalysts for chemoselective hydrogenation reactions via optimizing the parameters of the double-flame spray pyrolysis. *ACS Catal.* **2016**, *6*, 2372-2381.
- (46) Jiao, J.; Kanellopoulos, J.; Behera, B.; Jiang, Y. J.; Huang, J.; Marthala, V. R. R.; Ray, S. S.; Wang, W.; Hunger, M. Effects of adsorbate molecules on the quadrupolar interaction of framework aluminum atoms in dehydrated zeolite H, Na-Y. *J. Phys. Chem. B* **2006**, *110*, 13812-13818.
- (47) Wang, Q.; Li, W.; Hung, I.; Mentink-Vigier, F.; Wang, X.; Qi, G.; Wang, X.; Gan, Z.; Xu, J.; Deng, F. Mapping the oxygen

structure of γ -Al₂O₃ by high-field solid-state NMR spectroscopy. *Nat. Commun.* **2020**, *11*, 3620.

(48) Valla, M.; Rossini, A. J.; Caillot, M.; Chizallet, C.; Raybaud, P.; Digne, M.; Chaumonnot, A.; Lesage, A.; Emsley, L.; van Bokhoven, J. A.; Coperet, C. Atomic description of the interface between silica and alumina in aluminosilicates through dynamic nuclear polarization surface-enhanced NMR spectroscopy and First-Principles calculations. *J. Am. Chem. Soc.* **2015**, *137*, 10710-10719.

(49) Hunger, M.; Freude, D.; Pfeifer, H. Magic-angle spinning nuclear-magnetic-resonance studies of water-molecules adsorbed on Brønsted-acid and Lewis-acid sites in zeolites and amorphous silica-aluminas. *J. Chem. Soc. Faraday Trans.* **1991**, *87*, 657-662.

(50) Lang, S.; Benz, M.; Obenaus, U.; Himmelmann, R.; Hunger, M. Novel approach for the characterization of Lewis acidic solid catalysts by solid-state NMR spectroscopy. *ChemCatChem* **2016**, *8*, 2031-2036.

(51) Yin, F.; Blumenfeld, A. L.; Gruver, V.; Fripiat, J. J. NH₃ as a probe molecule for NMR and IR study of zeolite catalyst acidity. *J. Phys. Chem. B* **1997**, *101*, 1824-1830.

(52) Ma, D.; Han, X. W.; Xie, S. J.; Bao, X. H.; Hu, H. B.; Au-Yeung, S. C. F. An investigation of the roles of surface aluminum and acid sites in the zeolite MCM-22. *Chem. Eur. J.* **2002**, *8*, 162-170.

(53) Chizallet, C.; Raybaud, P. Acidity of amorphous silica-alumina: From coordination promotion of Lewis sites to proton transfer. *ChemPhysChem* **2010**, *11*, 105-108.

(54) Chizallet, C.; Raybaud, P. Density functional theory simulations of complex catalytic materials in reactive environments: Beyond the ideal surface at low coverage. *Catal. Sci. Technol.* **2014**, *4*, 2797-2813.

(55) Wang, Z.; Wang, L.; Jiang, Y.; Hunger, M.; Huang, J. Cooperativity of Brønsted and Lewis acid sites on zeolite for glycerol dehydration. *ACS Catal.* **2014**, *4*, 1144-1147.

(56) Wang, C.; Chu, Y.; Xu, J.; Wang, Q.; Qi, G.; Gao, P.; Zhou, X.; Deng, F. Extra-framework aluminum-assisted initial C-C bond formation in methanol-to-olefins conversion on zeolite H-ZSM-5. *Angew. Chem. Int. Ed.* **2018**, *57*, 10197-10201.

(57) Wang, Z.; O'Dell, L. A.; Zeng, X.; Liu, C.; Zhao, S.; Zhang, W.; Gaborieau, M.; Jiang, Y.; Huang, J. Insight into three-coordinate aluminum species on ethanol-to-olefin conversion over ZSM-5 zeolites. *Angew. Chem. Int. Ed.* **2019**, *58*, 18061-18068.

(58) Mojet, B. L.; Kappers, M. J.; Miller, J. T.; Koningsberger, D. C. Metal-support interactions in supported platinum catalysts: zeolites and amorphous supports. *Studies in Surf. Sci. and Catal.* **1996**, *101*, 1165-1174.

(59) Ma, L.; Cheng, Y.; Cavataio, G.; McCabe, R. W.; Fu, L.; Li, J. *In situ* DRIFTS and temperature-programmed technology study on NH₃-SCR of NO_x over Cu-SSZ-13 and Cu-SAPO-34 catalysts. *Appl. Catal. B-Environ.* **2014**, *156-157*, 428-437.

(60) Zhang, R.; Liu, N.; Lei, Z.; Chen, B. Selective transformation of various nitrogen-containing exhaust gases toward N₂ over zeolite catalysts. *Chem. Rev.* **2016**, *116*, 3658-3721.